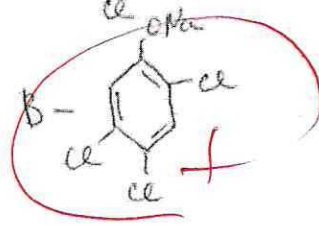
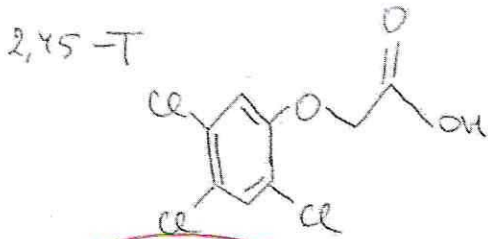
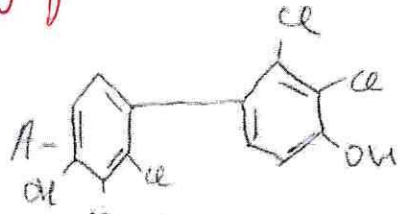
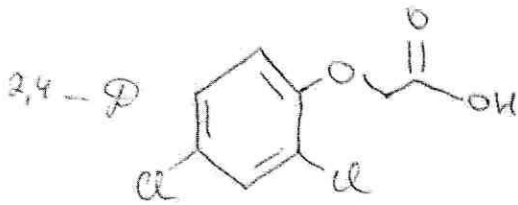


*Иванов*

1.1 Задача



$M_r(B) = 219.5$

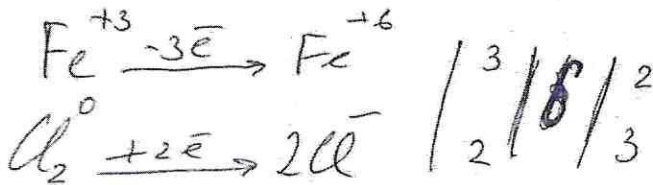
$M_r(A) = 324$

1	2	3	4	5	6	7	8	9	10
2		8	6	6	12	12			



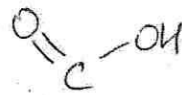
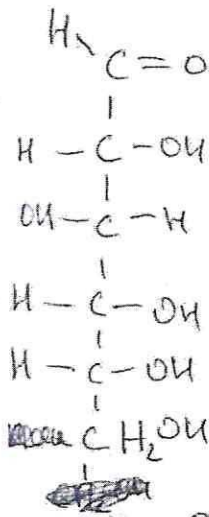
7.1 Задача

- 1)  $4\text{Fe}(\text{NO}_3)_2 \xrightarrow{t^\circ} 2\text{Fe}_2\text{O}_3 + 8\text{NO}_2 + \text{O}_2$  (X<sub>1</sub>) 2
- 2)  $\text{Fe}_2\text{O}_3 + 6\text{HBr} \rightarrow 2\text{FeBr}_3 + 3\text{H}_2\text{O}$  2
- 3)  $2\text{FeBr}_3 + 3\text{Na}_2\text{SO}_3 + 3\text{H}_2\text{O} \rightarrow 6\text{NaBr} + 2\text{Fe}(\text{OH})_3 + 3\text{SO}_2$  (X<sub>2</sub>) 2
- 4)  $\text{Fe}(\text{OH})_3 + 3\text{NaOH}_{\text{конц}} \rightarrow \text{Na}_3[\text{Fe}(\text{OH})_6]$  (X<sub>3</sub>) 2
- 5)  $2\text{Na}_3[\text{Fe}(\text{OH})_6] + 3\text{CO}_2 \rightarrow 3\text{Na}_2\text{CO}_3 + 2\text{Fe}(\text{OH})_3 + 3\text{H}_2\text{O}$  (X<sub>4</sub>) 2
- 6)  $2\text{Fe}(\text{OH})_3 + 3\text{Cl}_2 + 10\text{KOH} \rightarrow 6\text{KCl} + 2\text{K}_2\text{FeO}_4 + 8\text{H}_2\text{O}$  (X<sub>4</sub>) 2

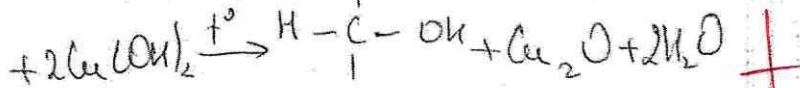


12

4.1 Задача



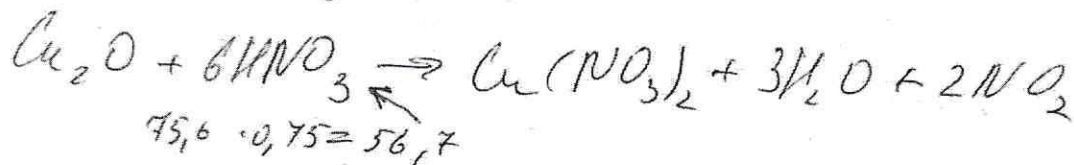
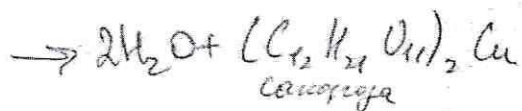
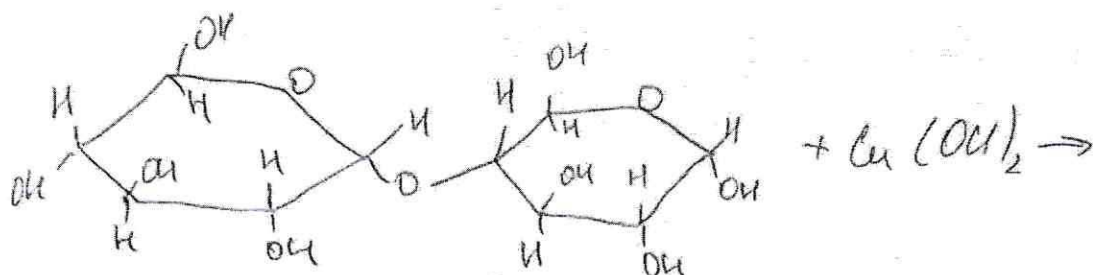
красно-фиолетовый оксид



СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ

118 147

4.1 продолжение



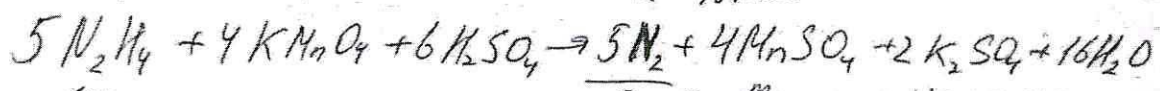
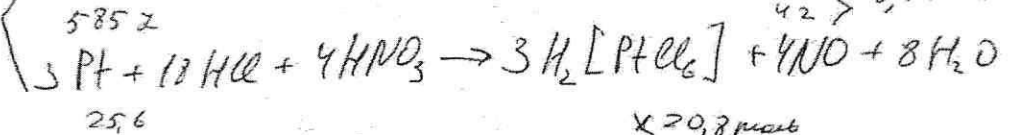
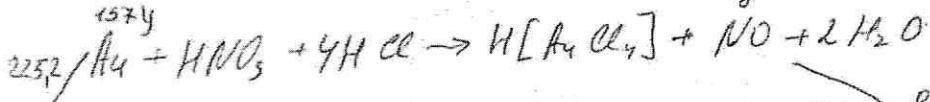
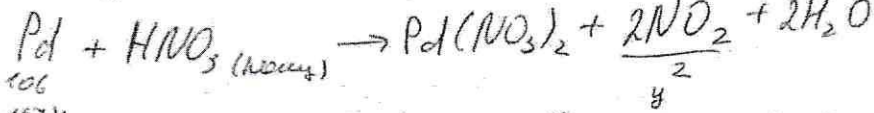
5.1 Задача

Pd, Pt, Au

$$V_{\text{использ}} = \frac{1}{3} \cdot \pi \cdot R^2 \cdot h = \frac{1}{3} \cdot 3,14 \cdot 1,2^2 \cdot 3 = 4,5216$$

12.800 → 12,8 г/мл    Δ радиус = 1,2 см  
h = 3 см

x = 353,616



0,8 · 8,34 = 0,098 моль

$\rho = \frac{m}{V}$      $m = \rho \cdot V = 12,8 \cdot 4,5216 = 57,88$

57,88 - 353,6 = ~~295,72~~ 15,48

107y + 585z = ~~295,72~~ 15,48

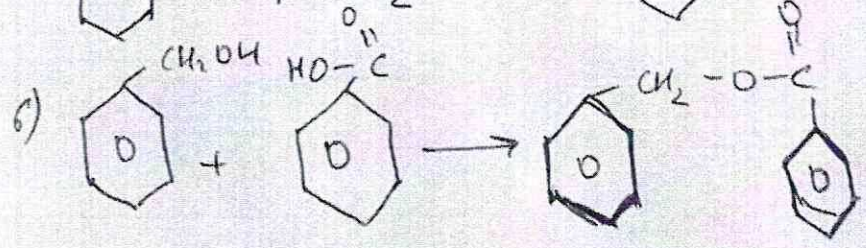
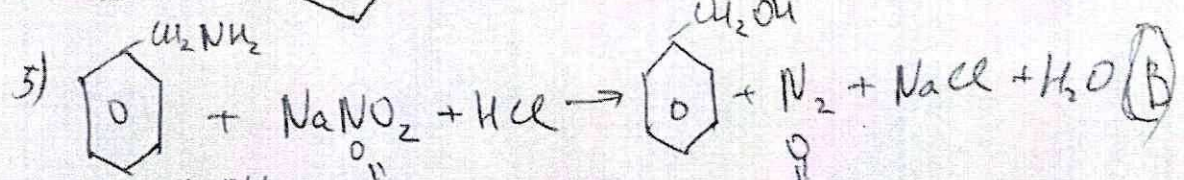
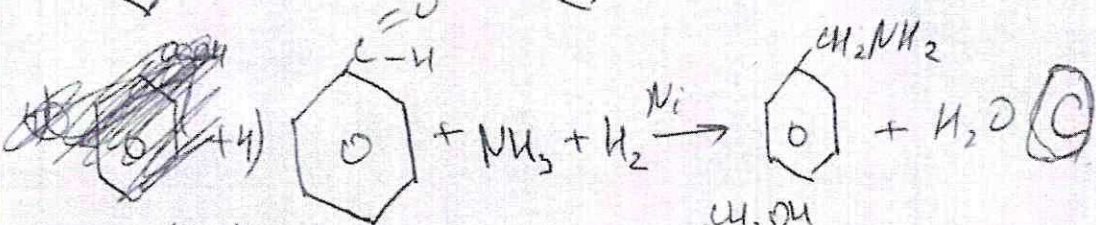
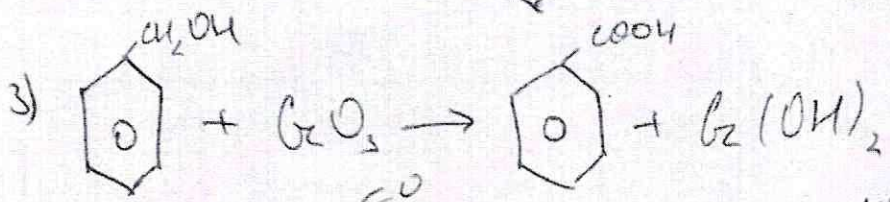
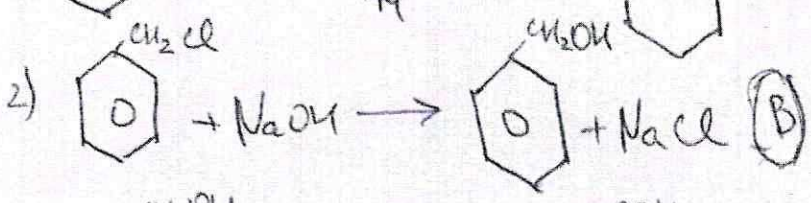
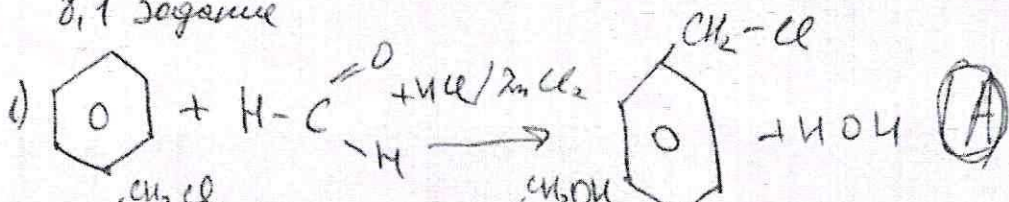
$W(\text{Pd}) = \frac{42,4 \cdot 100\%}{57,88} = 73,255\%$

y + 4z = 0,096    y = 0,0284

$W(\text{Pt}) = \frac{585 \cdot 0,017 \cdot 100\%}{57,88} = 17,182\%$

$W(\text{Au}) = \frac{197 \cdot 0,0284 \cdot 100\%}{57,88} = 9,666\%$

3.1 Задача



3.1 Задача  $V = h\pi r^2 = 1570$

$V' = 0,6 \cdot 1570 = 942$

$V(M_3) = 14,6 \quad n(M_3) = 0,652$

$c = 0,692 \text{ моль/л}$

$pH(M_3) = 14 - \frac{1}{2}(pK_1 - \lg c) = 11,5$

5 СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ

+  
12

+8