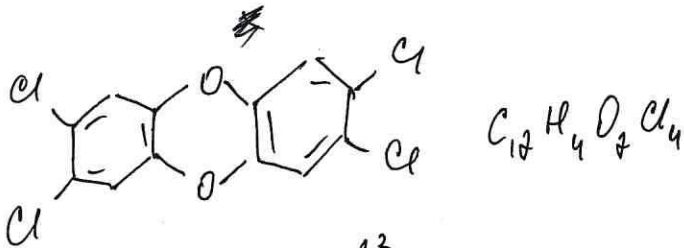


20 June

1	2	3	4	5	6	7	8	9	10
6	10	8	8	6		12	12		20

Задание ~ 1.2

A- тетра хлор бензо- пара- диоксин



$$n(C) = \frac{1,805 \cdot 10^{23}}{6,02 \cdot 10^{23}} = 0,3 \text{ моль}$$

$$n(Cl) = \frac{0,6627 \cdot 10^{23}}{6,02 \cdot 10^{23}} = 0,11 \text{ моль}$$

Пусть $n(A) = x$, а $n(B) = y$

$$\begin{cases} 12x + 6y = 0,3 \\ 4x + 3y = 0,11 \quad | \cdot 2 \end{cases}$$

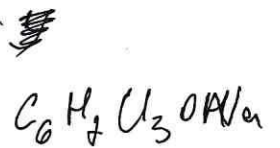
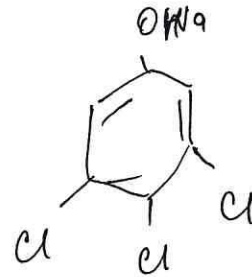
$$\begin{cases} 12x + 6y = 0,3 \\ - 8x + 6y = 0,22 \end{cases}$$

$$4x = 0,08$$

$$x = 0,02$$

$$y = 0,01$$

B- 2,4,5 трихлор фенол натрия



$$M(A) = 322$$

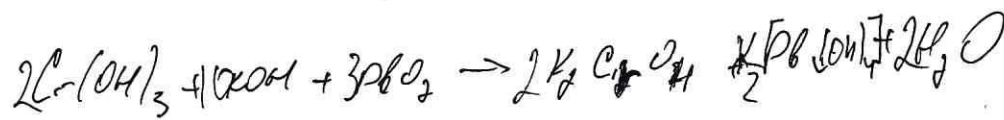
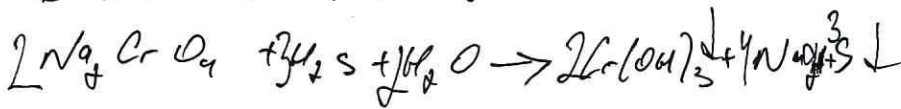
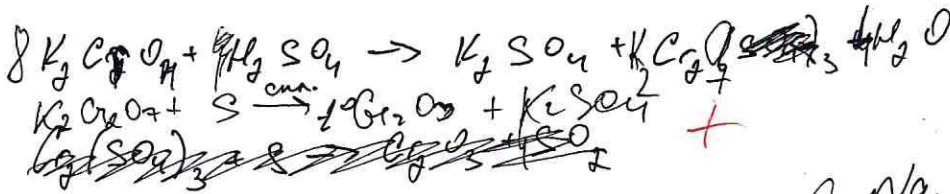
$$M(B) = 203,5$$

$$m_{\text{вещ}} = 322 \cdot 0,02 + 203,5 \cdot 0,01 = 6,44 + 2,035 = 8,475$$

Отв: 8,4752

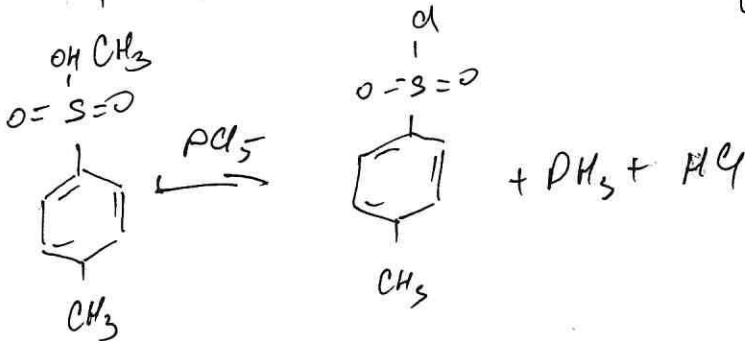
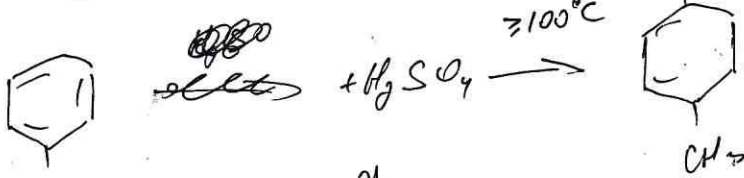


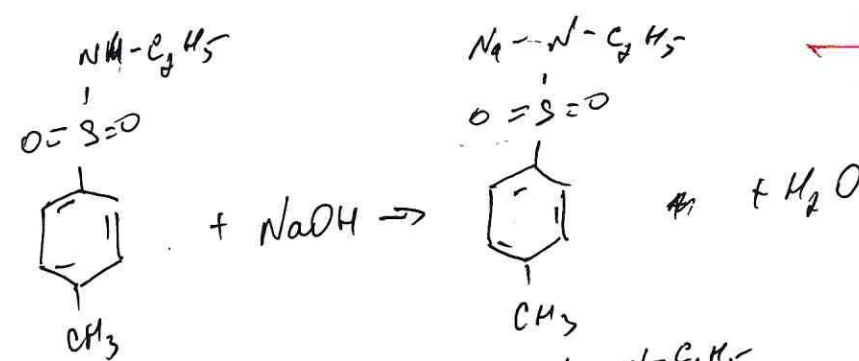
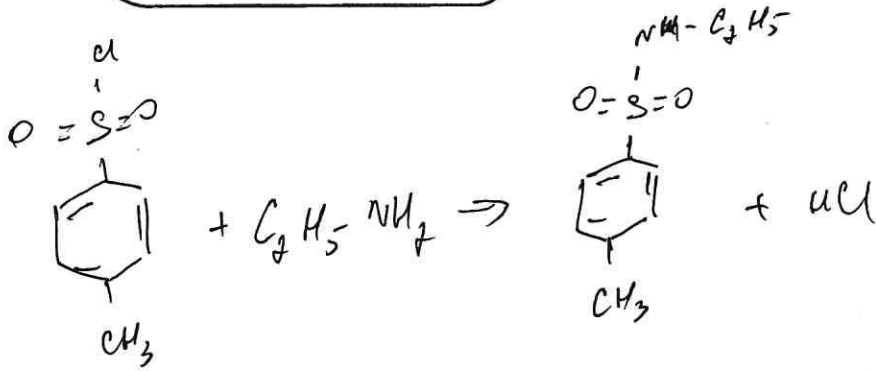
Задача 17.2



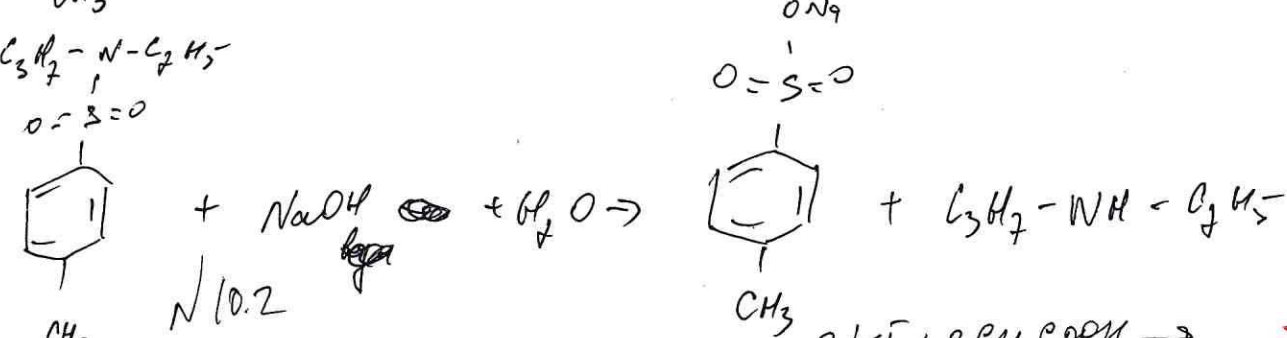
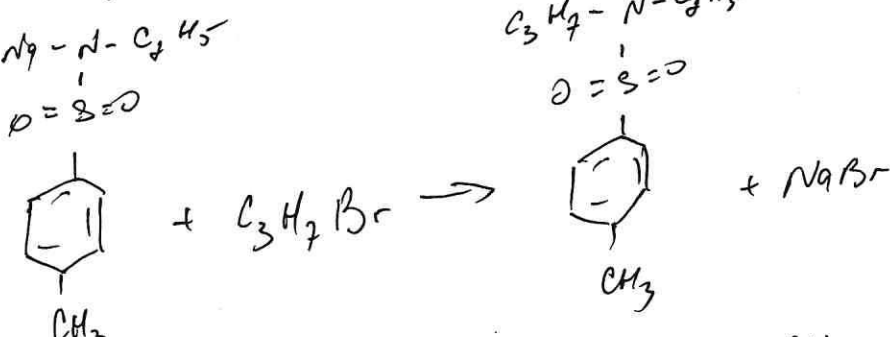
неактуально !!!
125

Задача 18.2

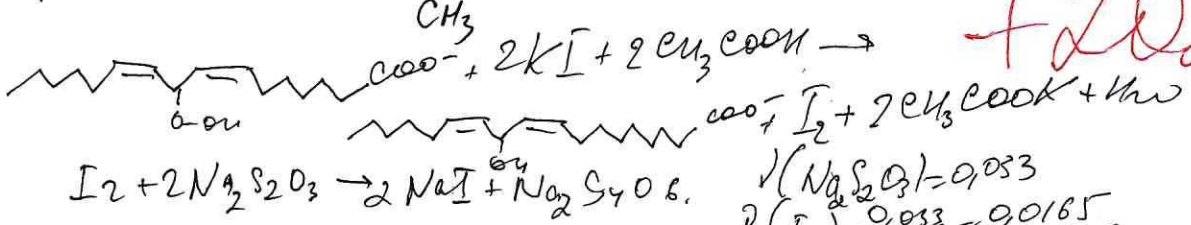
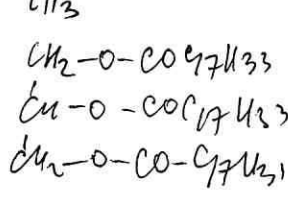




+ 120



+ 200



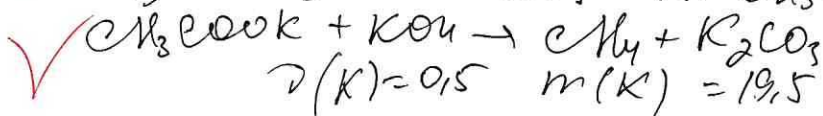
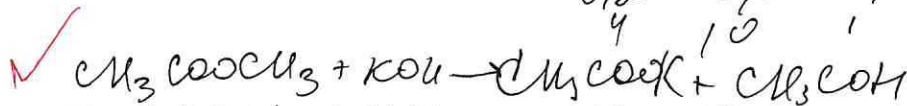
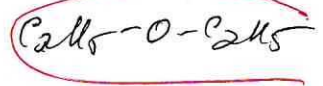
$\sqrt{(\text{Na}_2\text{S}_2\text{O}_3)} = 0,033$
 $\nu(\text{I}_2) = 0,033 = 0,0165$
 $\nu(\text{O}) = 0,0165 - 0,52$
 $\nu(\text{O}) = 3,17 - 6 \cdot 1000$
 $\text{I}_\nu = 3,17$ негор.



№ 2.2. $pV = \nu RT$ $\nu = 0,2$

$\nu(C) = \nu(CO_2) = 0,2$; $\nu(H_2O) = 0,25$ $\nu(H) = 0,5$ $\nu(O) = 0,05$

$C: H: O$
 $0,2 : 0,5 : 0,05$
 $4 : 10 : 1$



$\nu(K) = 0,15$ $m(K) = 19,5$
 $m(C_4H_{10}O) = 44 \cdot 0,2 = 14,8$
 $m(C_4H_8 + C_4H_9OH) = 19,2$

$\omega = 41\%$

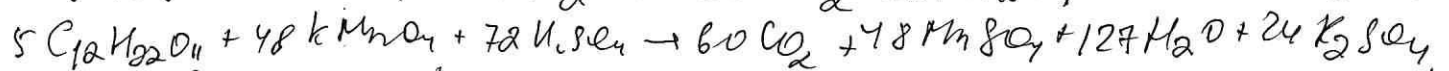
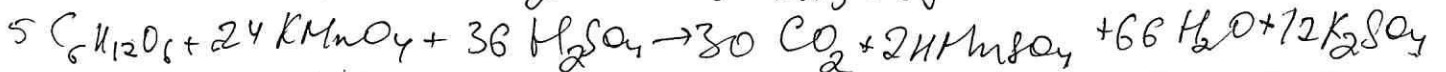
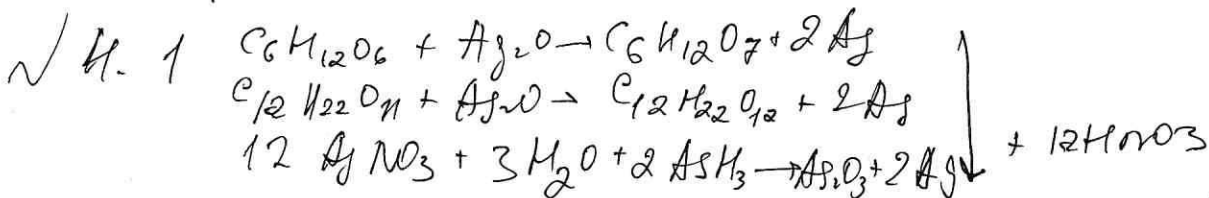
+

№ 3.1 $V = h \cdot \pi \cdot r^2 = 490,625$ $V' = 0,6 \cdot 490,625 = 294,375$

$\nu(C_4H_8O) = 0,623$ $C = \frac{0,623}{0,294375} = 2,11$

$PH = 0,5(pKa - \lg C(C_4H_8O)) = 2,2$

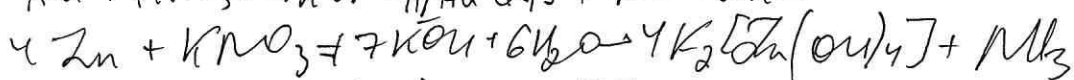
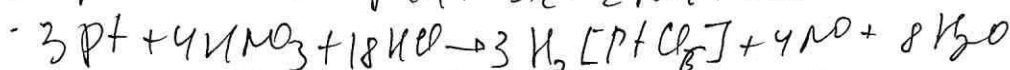
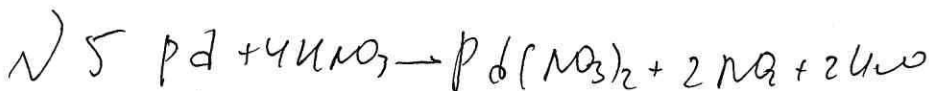
85



$\nu(AgNO_3) = \frac{m}{M} = 0,13$ $x + y = 0,15$ $x = 0,1$
 $\nu(CO_2) = \frac{pV}{RT} = 1,2$ $6x + 12y = 1,2$ $y = 0,05$

$m_{сум} = 35,1$

$C_6H_{12}O_6 + [C] \rightarrow C_6H_{14}O_6$ $m(C_6H_{14}O_6) = 0,1 \cdot 182 (100\%)$
 $14,56 (80\%)$



$\nu(Zn) = 3,2$ $\nu(NH_3) = 0,8$ $V_{H_2} = \pi R h = 4,526$ $m = 57,9$

$m(Pd) = 42,4$ $\omega(Pd) = 73\%$
 $m(Pt) = 9,8$ $\omega(Pt) = 17\%$
 $m(Au) = 5,7$ $\omega(Au) = 10\%$

