

1	2	3	4	5	6	7	8	9	10
8	10	8	6	10	1	1	4	2	5

Σ 55

№ 1.2.

$$6N(O) = N(C)$$

$$4N(O) = N(H) + N(Cl)$$

$$N(H) = N(Cl) \text{ по услов.}$$

$$\Rightarrow 4N(O) = 2N(H)$$

$$2N(O) = N(H)$$

$$22 = N(H) + N(O) + N(Cl) + N(C)$$

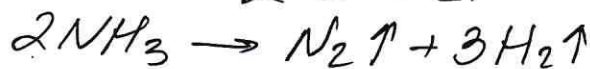
$$22 = 6N(O) + N(O) + 2N(O)$$

$$22 = 9N(O)$$

$$N(O) = 2 \quad N(C) = 12 \quad N(Cl) = 4 \quad N(H) = 4$$

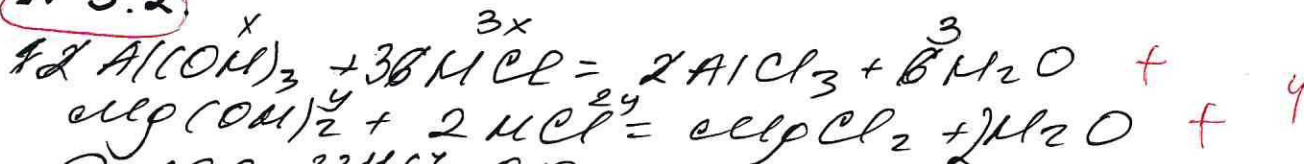


№ 2.2.



10

№ 3.2



$$Q_{HCl} = \frac{2311,672 \cdot 0,15}{36,5 \text{ г/моль}} = 9,497 \text{ моль}$$

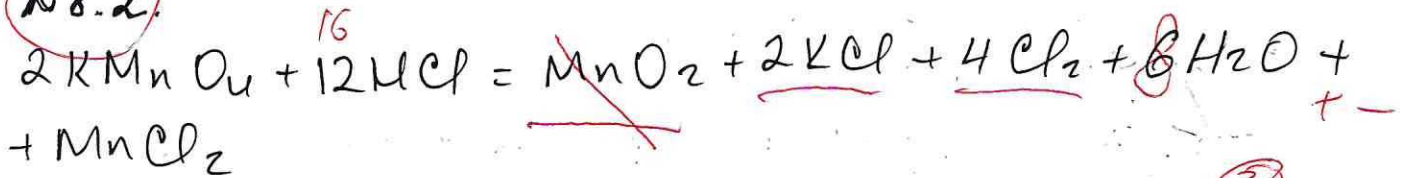
$$\begin{cases} 78x + 58y = 262 \\ 3x + 2y = 9,497 \end{cases}$$

СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ

8 x 0 7 2

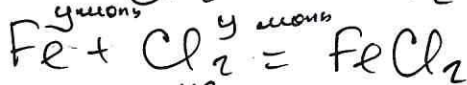
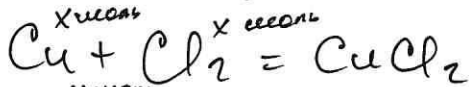


№ 8.2.



$$n_{\text{KMnO}_4} = \frac{3,162}{158 \text{ г/моль}} = 0,02 \text{ моль} \Rightarrow +$$

$$\Rightarrow n_{\text{Cl}_2} = 0,04 \text{ моль}$$



$$\begin{cases} 64x + 56y = 3 \\ x + 2y = 0,04 \quad | \cdot 64 \end{cases}$$

$$80y = 0,44$$

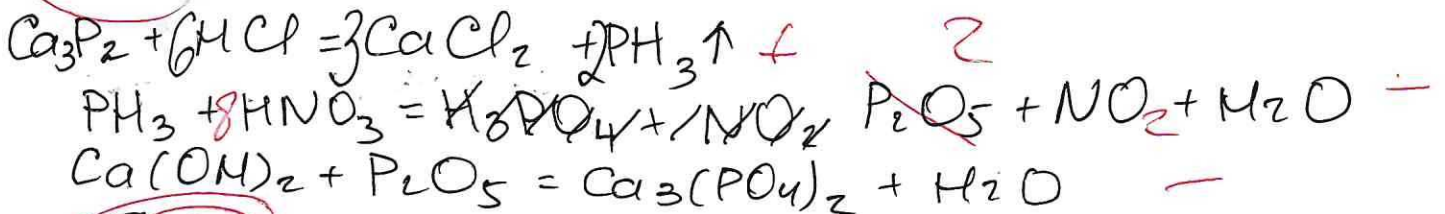
$$\sum y = 0,0055 \text{ моль}$$

$$\sum x = 0,035 \text{ моль} \quad 0,0235 \text{ моль}$$

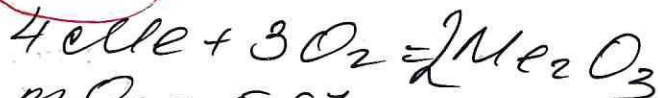
$$m_{\text{Cu}} = 1,5042$$

$$\omega(\text{Cu}) = 50,13\% \quad (-)$$

№ 9.2.



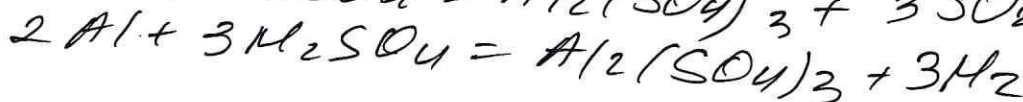
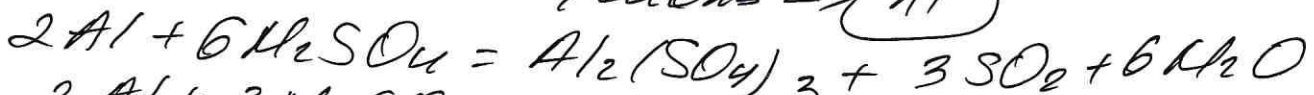
№ 10.2.



$$m_{\text{O}_2} = 5,672 - 32 = 2,672$$

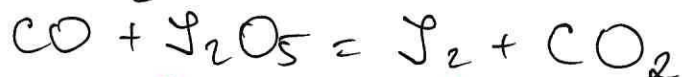
$$n(\text{O}_2) = 0,0834 \text{ моль} \Rightarrow n_{\text{Me}} = 0,1251 \text{ моль}$$

$$m(\text{Me}) = 27 \cdot 0,1251 = 3,3777 \text{ г} \Rightarrow \text{Al}$$



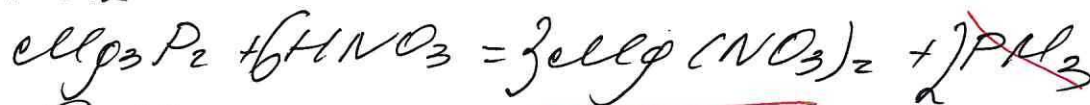


№ 4.2.



7

№ 6.2



$$D_{eP_2} = 0, \text{ eмоль} \Rightarrow$$

$$\Rightarrow D_{P_2O_5} = 0,2 \text{ моль}$$

$$m_{P_2O_5} = 6,82$$

1

