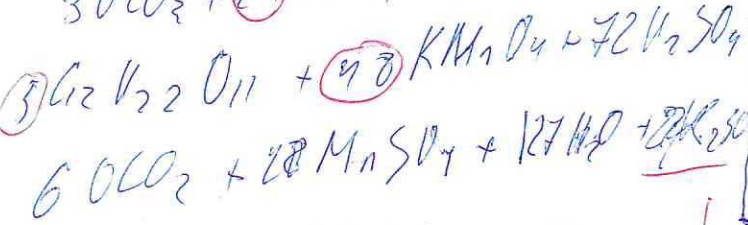
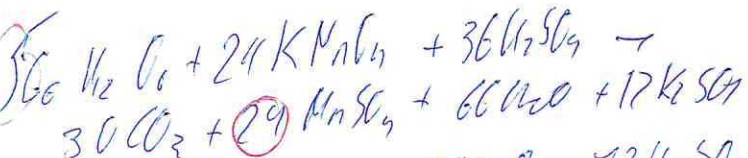
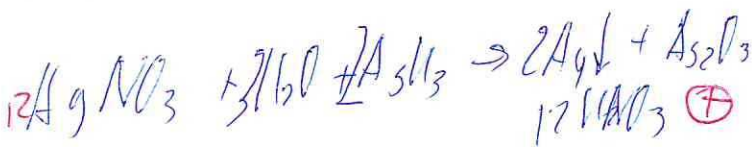
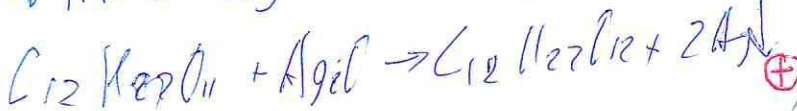
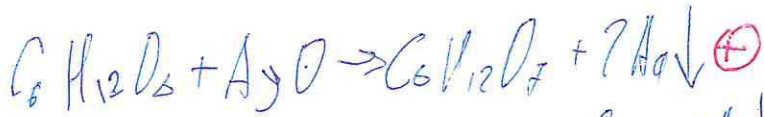


№4.2

№4.2

№7.2



$$V(AgNO_3) = \frac{510 \cdot 0,1}{770} = 0,3 \oplus$$

$$V(Ag) = V(AgNO_3) = 0,3$$

количество x $V(Ag) = 2x$
масса X $V(Ag) = 2x$

$$V(CO_2) = \frac{101,3 \cdot 97,8}{2,31 \cdot 223} = 1,2 \oplus$$

$$V(CO_2)_1 = 6x$$

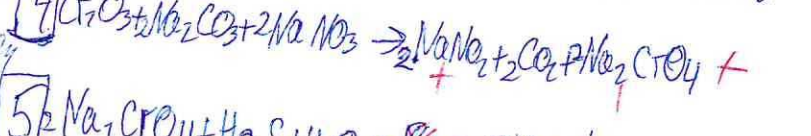
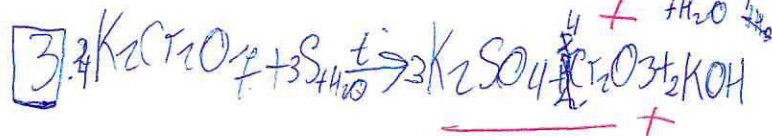
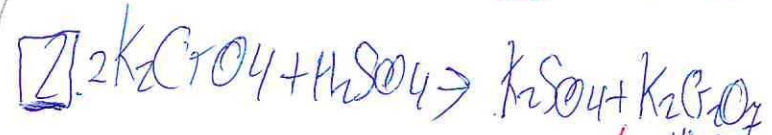
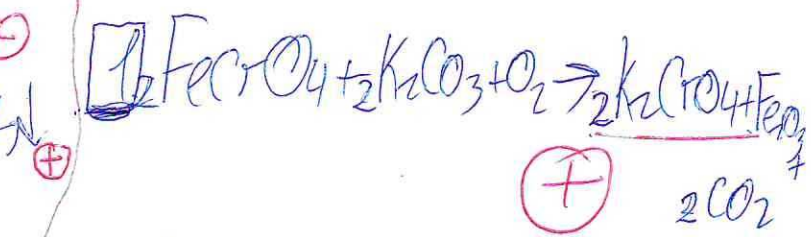
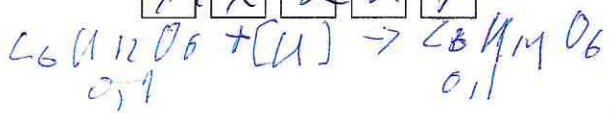
$$V(CO_2)_2 = 12x$$

$$m(CO_2) = 35,1 \oplus$$



СЕЧЕНОВСКИЙ
УНИВЕРСИТЕТ

А А А А А



6 ?

1	2	3	4	5	6	7	8	9	10
5	9	7	7	6	5	10	12	7	47

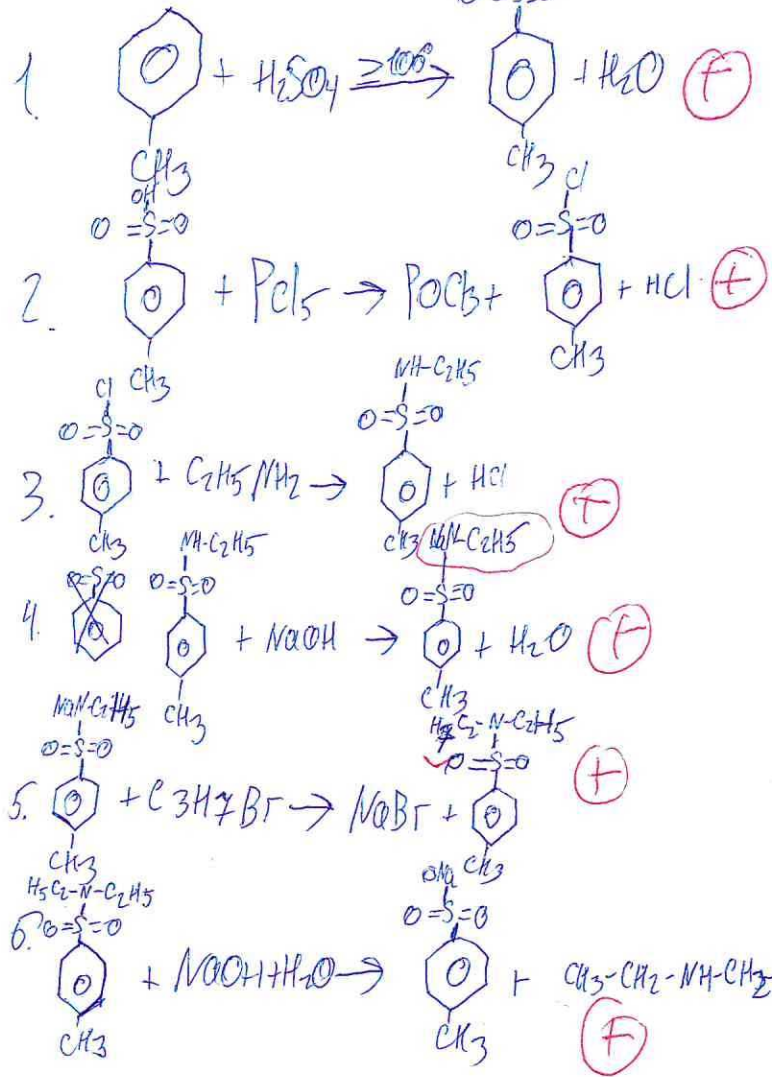
$$\left. \begin{aligned} x + y &= 0,15 \\ 6x + 12y &= 1,2 \end{aligned} \right\} \begin{aligned} x &= 0,1 \\ y &= 0,05 \end{aligned}$$

Σ 85 г

$$m(C_6H_{14}O_6) = 0,1 \cdot 182 = 18,2$$

$$12,2 \cdot 0,78 = 9,516 = 80\% \oplus$$

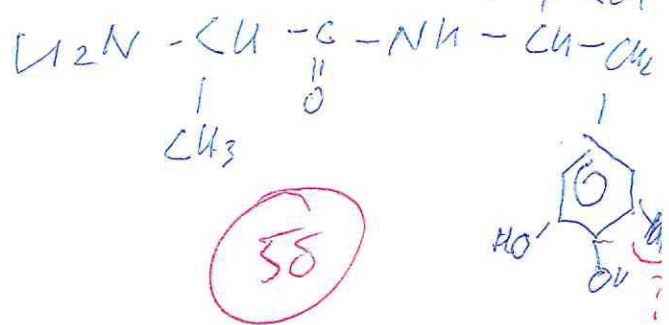
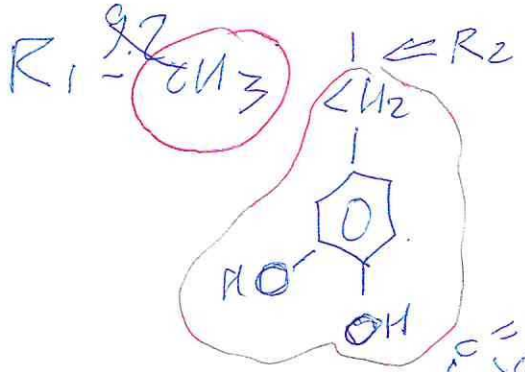
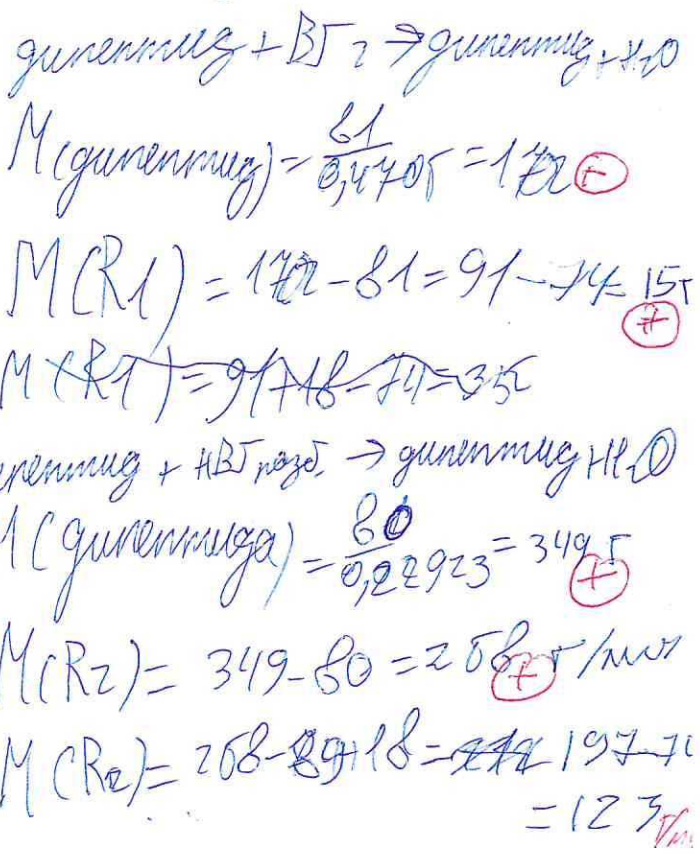
№ 2.



(12)

62

б. 2.



□ □ □ □ □

2.

$$101 \cdot 4,82 = D = 8,31 \cdot 243k$$

$$D = 0,2 \oplus$$

$$D(\text{CO}_2) = 0,2 \cdot 12 = 2,4 \oplus$$

$$D(\text{H}_2\text{O}) = \frac{95}{12} = 0,25 \oplus$$

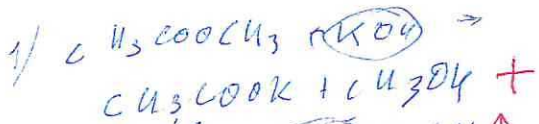
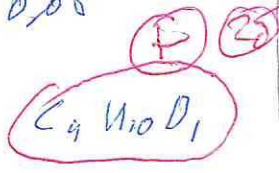
$$D(\text{H}_2\text{O}) = 0,25 \cdot 2 = 0,5 \oplus$$

$$m(\text{H}) = 0,5 \cdot 1 = 0,5$$

$$m(\text{O}) = 3,7 - 2,4 - 0,5 = 0,8$$

$$D(\text{O}) = \frac{0,8}{16} = 0,05 \oplus$$

x	y	z
C	H	O
4	10	1



$$D(\text{KOH}) = 0,125 \cdot 4 = 0,5$$

$$D(\text{KOH}) = 0,5 - 2x$$

$$m_{\text{CO}_2} = m(\text{K}_2\text{CO}_3) - m(\text{KOH}) = 138x - 56(0,5 - 2x)$$

$$m_{\text{CO}_2} = 26x + 28$$

$$D = 0,5 | \text{K} | \quad m(\text{K}) = 19,5$$

$$19,5 = 15,77 \cdot x \quad 15,77 \quad x = 0,2$$

$$m(\text{группа}) = 10,2$$

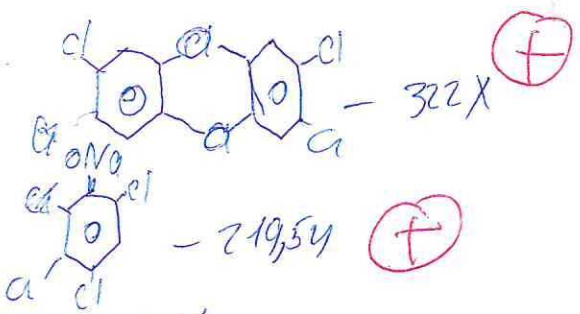
$$m(\text{C}_4\text{H}_8\text{OOCCH}_3) = 74 \cdot 0,2 = 14,8$$

СЕЧЕНОВСКИЙ
УНИВЕРСИТЕТ

1
 2
 3
 4
 5

95

1.2.



$$m = \frac{322x}{322x + 219,54}$$

$$1) D(\text{C}) = \frac{1,805 \cdot 10^{23}}{8,02 \cdot 10^{23}} = 0,3 \oplus$$

$$2) D(\text{Cl}) = \frac{0,8822 \cdot 10^{23}}{8,02 \cdot 10^{23}} = 0,11 \oplus$$

$$\begin{cases} 4x + 3y = 0,11 \\ 12x + 6y = 0,3 \end{cases}$$

$$1) 6y = 0,3 - 12x$$

$$2) y = 0,05 - 2x$$

$$3) 4x + 0,15 - 6x = 0,11$$

$$4) -2x = 0,04$$

$$5) x = 0,02$$

$$\begin{cases} 6) 0,24 + 6y = 0,3 \\ 7) 6y = 0,06 \\ 8) y = 0,01 \end{cases}$$

$$x = 0,02 \quad y = 0,01$$

$$m = \frac{322 \cdot 0,02}{322 \cdot 0,02 + 219,54 \cdot 0,01} = 6,44 \quad | \quad \text{C}_{12}\text{H}_4\text{Cl}_{14}\text{O}_2$$

$$m = \frac{322 \cdot 0,02 + 219,54 \cdot 0,01}{322 \cdot 0,02 + 219,54 \cdot 0,01} = 2,195 \quad | \quad \text{C}_6\text{H}_2\text{Cl}_3\text{O}_1$$

$$6,44 = 6,44 \cdot 2,195$$

$$6,44 = 8,835 \quad m(6,44 + 2,195) = 8,635 \oplus$$

$$m = 0,744 \quad \Sigma$$

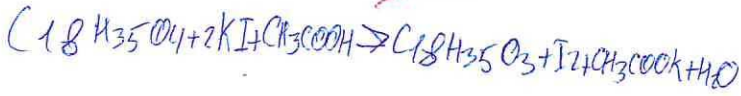
ω(группа) = 49,8%

+

65

10.7.

Структур формула?



1. $V_0 - V_k = 3,5 - 0,2 = 3,3 \quad +$

2) $\nu(Na_2S_2O_3) = 3,3 \cdot 0,01 = 0,033 \text{ (моль)} \quad +$

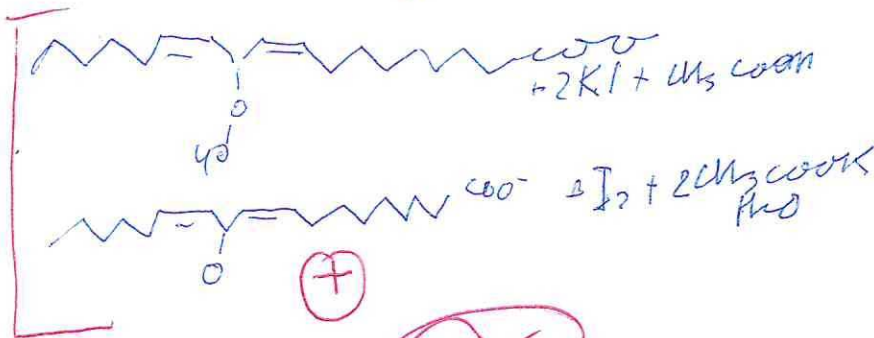
3) $\nu(I_2) = 0,033 : 2 = 0,0165 \text{ (моль)}$

4) $V(O) = 0,0165 \cdot 5,2 = 0,858 \text{ (г)}$

5) $\nu(O) = 0,0165 \cdot 1000 / 5,2 = 3,17$

6) $I_2 = 3,17$ - перекисл. число

Поскольку перекисл. число не должно превышать 3,5, то перекисл. число неадекватно



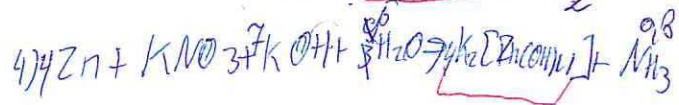
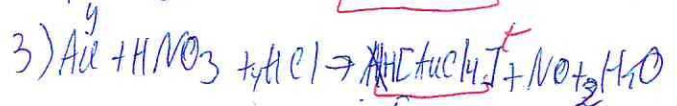
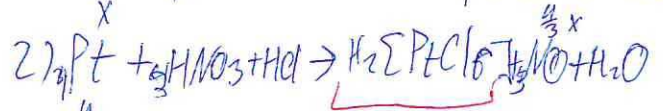
176



СЕЧЕНОВСКИЙ
УНИВЕРСИТЕТ

□ □ □ □ □

5.2



5) $V = \sqrt{R \cdot H} = \sqrt{3,11 \cdot 0,5^2} = 1,25 \quad +$
 $m(\text{соль}) = 4,52 - 12800 / 1000 = 5,79 \quad +$

6) $\nu(Zn) = \frac{208}{65} = 3,2 \quad +$

7) $\nu(NH_3) = 0,8 \quad +$

8) $\nu(NO_2) = 0,8$

9) $\nu(Pb) = 0,4$

10) $m(Pb) = 0,4 \cdot 100 = 42,4 \quad +$

$m(Au, Pt) = 5,79 - 4,52 = 1,27$
 $\nu(NO) = 0,0145 \cdot 95,9$

$\begin{cases} 195x + 197y = 1,55 \\ \frac{4}{3}x + y = 0,959 \end{cases}$

11) $4 = 0,959 \cdot \frac{4}{3}x$
 $x = 0,05 \quad y = 0,909$

11) $m(Au) = 0,028 \cdot 197 = 5,5$

12) $m(Pt) = 0,05 \cdot 195 = 9,75$

13) $w(Pt) = \frac{9,75}{57,9} = 0,168 = 16,8\%$

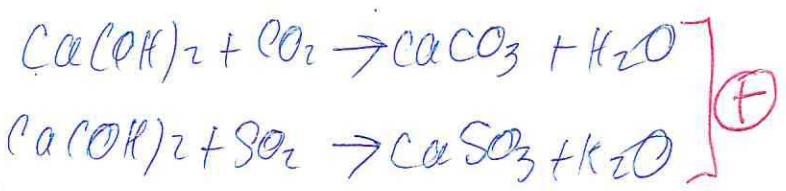
14) $w(Au) = \frac{5,5}{57,9} = 0,095 = 9,5\%$

15) $w(Pt) = \frac{9,75}{57,9} = 0,168 = 16,8\%$

6

$C_3H_7NO_2$ (9?)
Цемент - 111 ✓

Цимент - 123 ✗
 $C_3H_7N_3O_5$



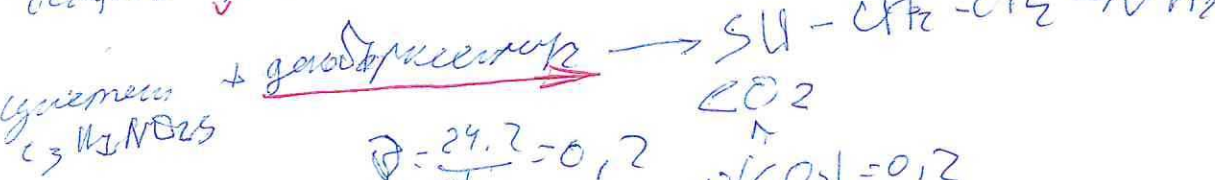
Цемент.
 $\nu(N) = 1/2 \nu(C) = x/2$
 $\nu(N_2) = 3x/2$ (цм) $3y/2$

$\frac{0,5x}{1,5y} = 2$
 $x = 6y$
 $m(CaCO_3) = 3x \cdot 100 = 1800y$
 $m(CaSO_3) = 720y = 120x$

1 осадок
 $\nu(CaCO_3) = 3x$
 $\nu(CaSO_3) = x$

2 осадок
 $\nu(CO_2) = \nu(CaCO_3) = 3y$

$\frac{0,09 \cdot 44}{0,09 \cdot 44} = \frac{2520y}{900y} = 2,8$



$V = 0,2 \cdot 22,4 = 4,48 \text{ л}$

(3.2)

1. $V = \pi R \cdot H = 3,14 \cdot 25^2 \cdot 10 = 19637,5$

2) $V(CO_2) = 19637,5 \cdot 0,05 = 981,875$

3) $V(CO_2) = 981,875$

$\nu(N)_{\text{кисл}} = 12,46 - 0,05 = 0,623 \text{ м}$

$C = \frac{\nu}{V} = \frac{0,623}{0,2943 \text{ л}} = 2,11$

$PH = 0,5 / \sqrt{K_a} - \log 2,11$

$= 0,5 / \sqrt{4,76} - \log 2,11 = 2,2$

(7)

(75)



□ □ □ □ □