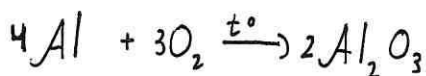


N	1	2	3	4	5	6	7	8	9	10
	4	5	-	10	8	2	-	-	6	10

45  
/

105

10.2



$$M(Al) = 27$$

$$M(Al_2O_3) = 102 \text{ г/моль}$$

$$n(Al) = \frac{3}{27} = 0,111 \text{ моль}$$

$$n(Al_2O_3) = n(Al) : 2; \quad n(Al_2O_3) = \frac{0,111}{2} = 0,0555 \text{ моль}$$

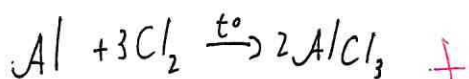
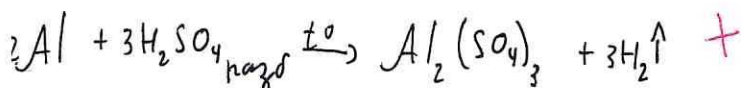
$$m(Al_2O_3) = 5,67 \text{ г}$$

$$M(Al_2O_3) = 102 \text{ г/моль}$$

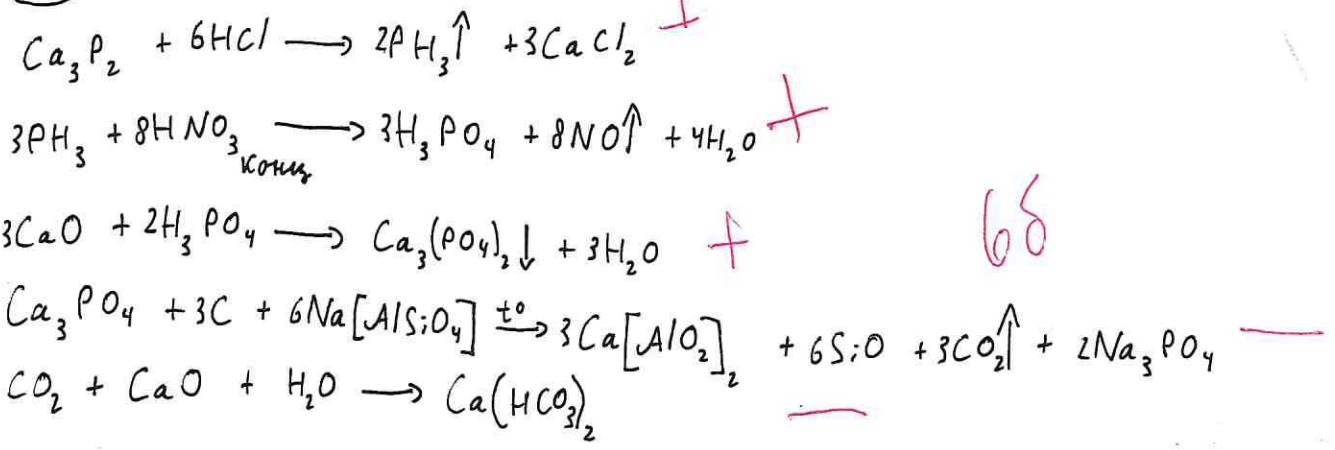
$$n(Al_2O_3) = \frac{5,67}{102} = 0,0555 \text{ моль}$$



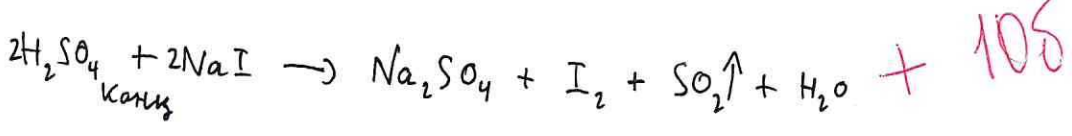
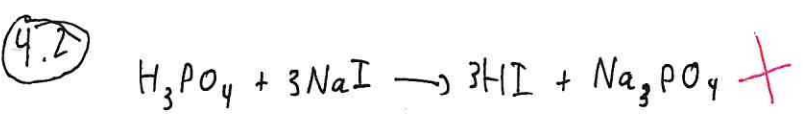
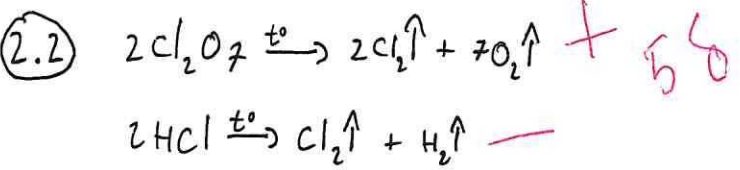
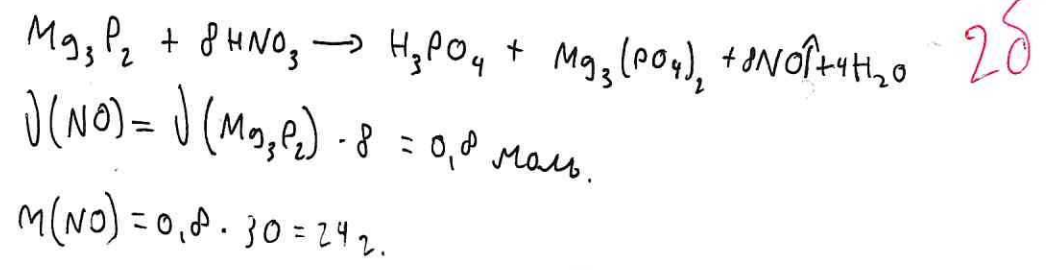
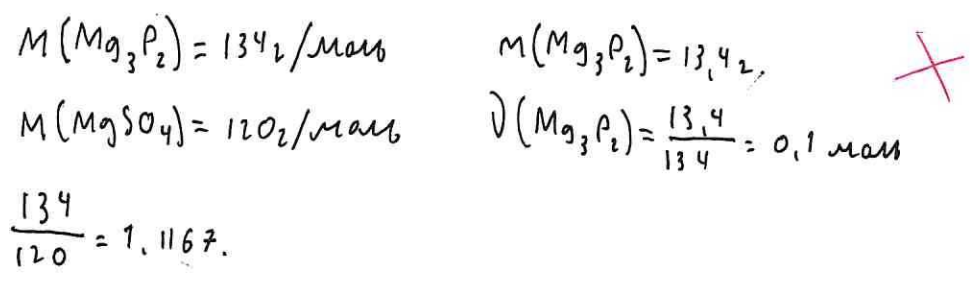
$Al + H_2SO_4 \text{ конц.} \rightarrow$  идет пассивация реакции +



9.2



6.2



СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ

$$2\text{I}^{-1} \xrightarrow{-2e^-} \text{I}_2^0 \quad | \quad -1$$

$$\text{S}^{+6} \xrightarrow{+2e^-} \text{S}^{+4} \quad | \quad +1$$

$\text{I}^{-1} (\text{NaI})$  - восстановитель  
 $\text{S}^{+6} (\text{H}_2\text{SO}_4)$  - окислитель

1.2

1-формула вещества  $C_{12}H_4Cl_4O_2$ .

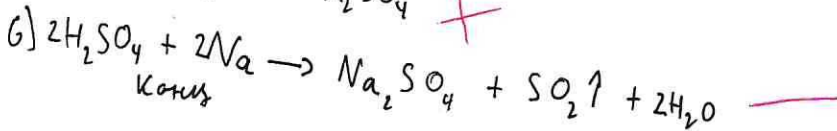
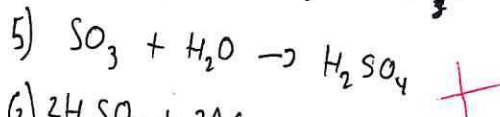
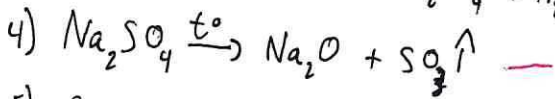
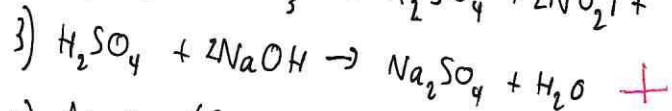
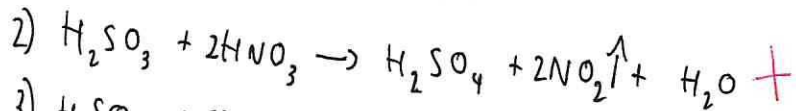
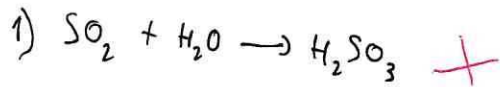
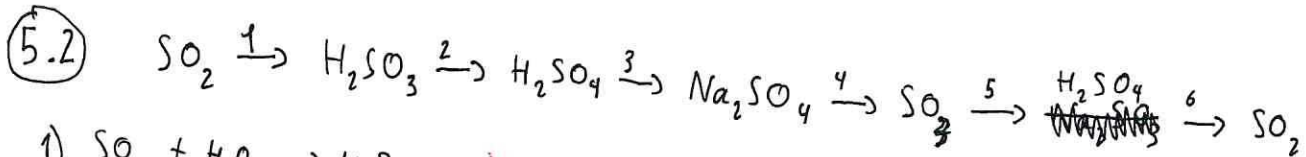
$$12 + 4 + 4 + 2 = 22$$

$$4 = 4$$

$$\frac{12}{6} = 2$$

$$\frac{4+4}{2} = 4$$

46



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