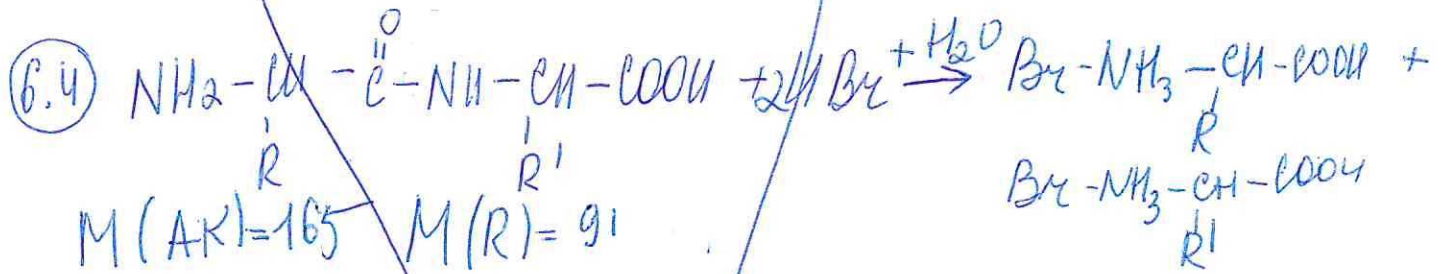


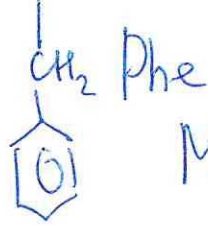
808 April

1	2	3	4	5	6	7	8	9	10
10	8	2	6	10	12	12		20	

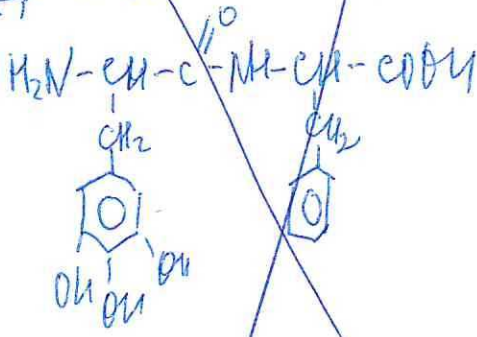


$M(\text{AK}) = 165$

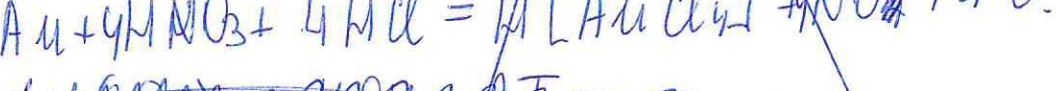
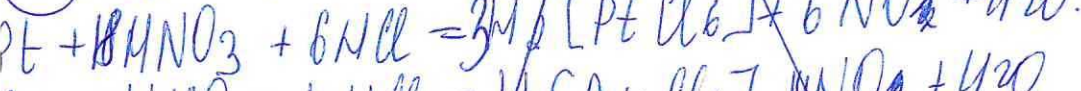
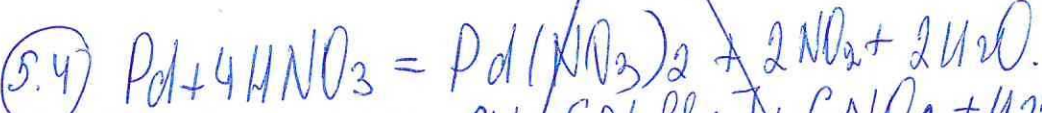
$M(\text{R}) = 91$



$M(\text{Д.Н.В.Т.}) = 441$   $M(\text{A}) = 360$   $M(\text{AK}) = 360 + 18 + 165 = 213$



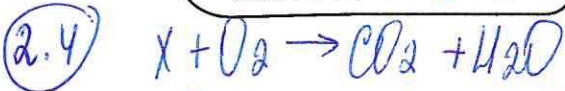
105



~~1)  $V(\text{сплав}) = \frac{4}{3} \pi r^3 = \frac{4}{3} \pi \cdot 1,02^3 = 4,4 \text{ см}^3$~~

~~2)  $m(\text{сплав}) = 13,032 \cdot 4,4 = 57,342$~~

3)



1)  $pV = \nu RT \Rightarrow 99,9 \cdot 14,72 = 8,31 \cdot 295 \Rightarrow \nu(O_2) = \frac{99,9 \cdot 14,72}{8,31 \cdot 295} = 0,6 \text{ моль}$

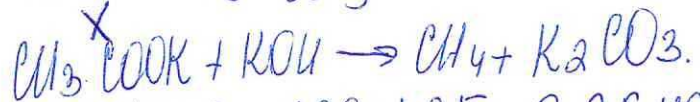
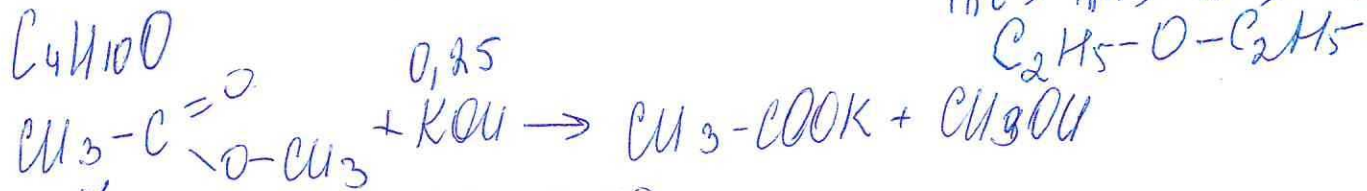
2)  $\nu(H_2O) = \frac{13,5}{18} = 0,75 \text{ моль}$

3)  $\nu(CO_2) = 0,6 \text{ моль} \Rightarrow \nu(C) = 0,6 \text{ моль} \Rightarrow m(C) = 0,6 \cdot 12 = 7,2$

$\nu(H_2O) = 0,75 \text{ моль} \Rightarrow \nu(H) = 1,5 \text{ моль} \Rightarrow m(H) = 1,5$

$\Rightarrow m(O_2) = 11,1 - 7,2 - 1,5 = 2,4$

4)  $\frac{7,2}{12} : \frac{1,5}{1} : \frac{2,4}{16} = 0,6 : 1,5 : 0,15 \Rightarrow 4 : 10 : 1$



5)  $\nu(KOH) = \frac{200 \cdot 1,25}{1000} = 0,25 \text{ моль} \Rightarrow m(KOH) = 142 \Rightarrow m(K) = 0,25 \cdot 39 = 9,75$

6)  $m(K) = \frac{9,75 \cdot 58,43}{100} = 5,722$   
 $\nu(CH_3COOCH_3) = x$   
 $\nu(KOH) = 2x$   
 $\nu(K_2CO_3) = x$

7)  $m(CH_3COOCH_3) = 7,4$   
 $m(C_2H_5OC_2H_5) = 7,6$   
 $\omega = 51\%$

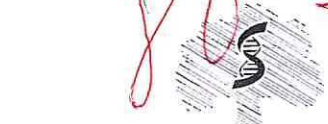
3.4) 1)  $V_{цилиндра} = \pi r^2 h = 3,14 \cdot 3,5^2 \cdot 30 = 1153,95 \text{ см}^3$   
 2)  $p = 1,0666 \text{ г/см}^3$ ;  $V = 50 \text{ см}^3 \Rightarrow m(CH_3COOH) = pV = 53,33$

3)  $m(CH_3COOH) = \frac{65 \cdot 5,33}{100} = 34,662 \Rightarrow \nu(CH_3COOH) = \frac{34,662}{60} = 0,5777 \text{ моль}$

4)  $\nu(CH_3COOH) = 0,5777$

5)  $\nu(H_2O) = \frac{70 \cdot 1153,95}{100} = 807,765 \text{ см}^3$

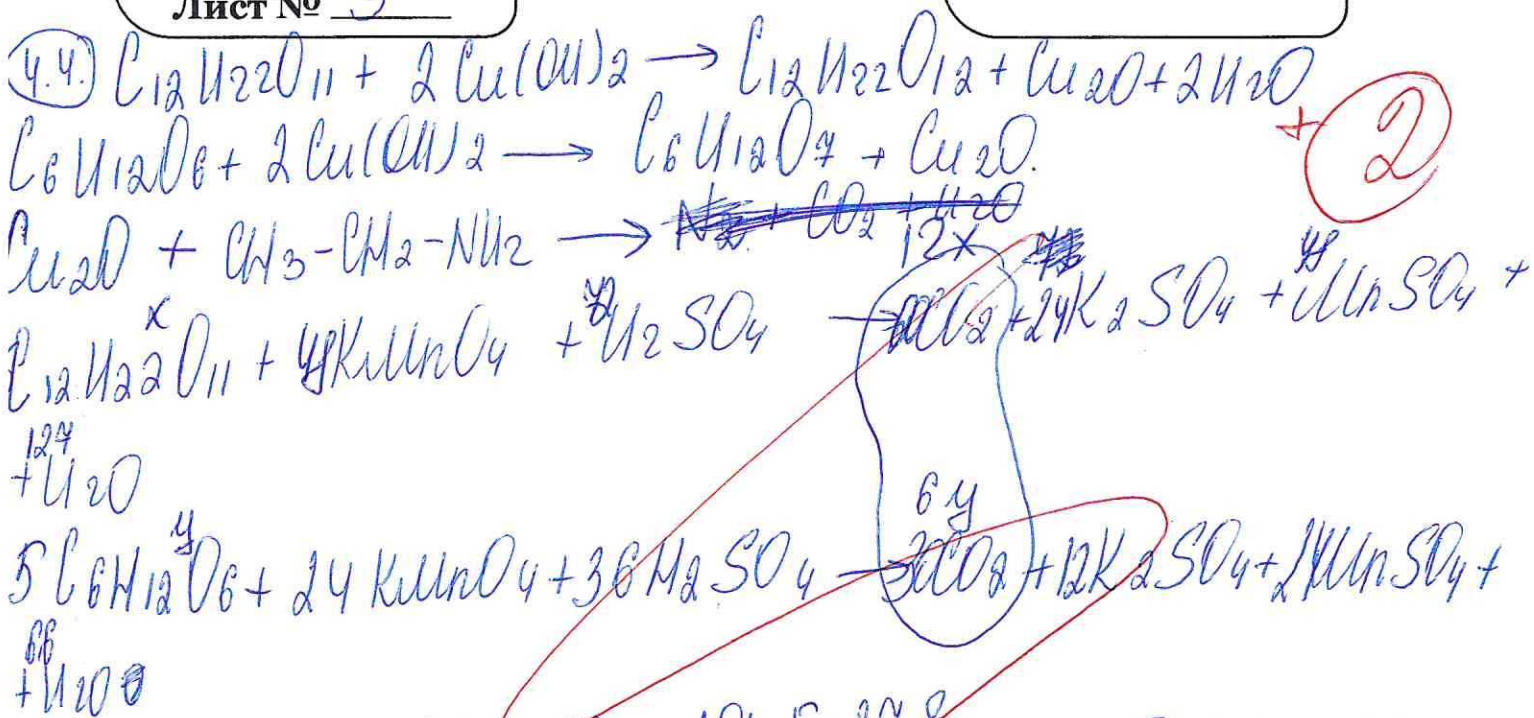
$pH(CH_3COOH) = \frac{1}{2}(pK_a - \lg c(CH_3COOH)) = 2,45$



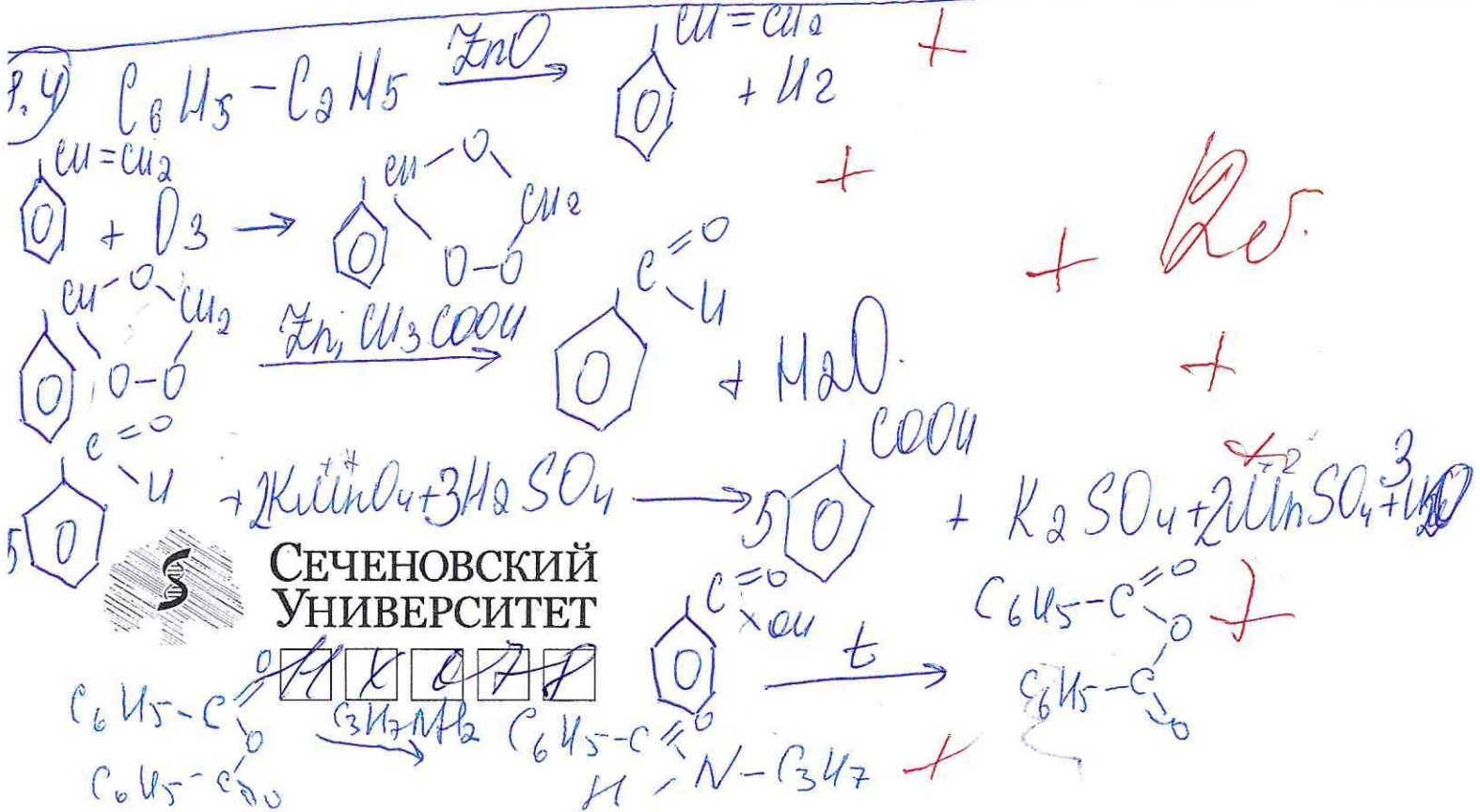
СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ



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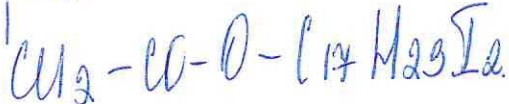
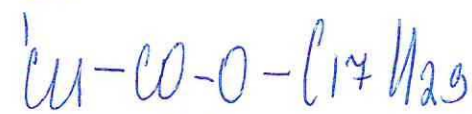
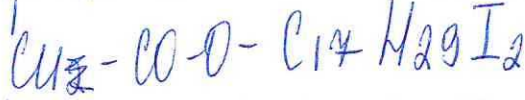
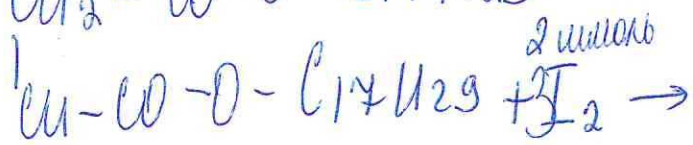
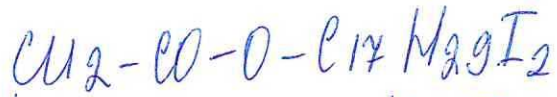
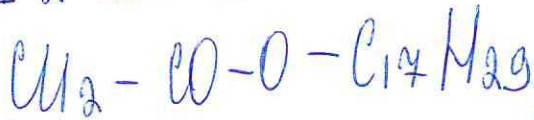
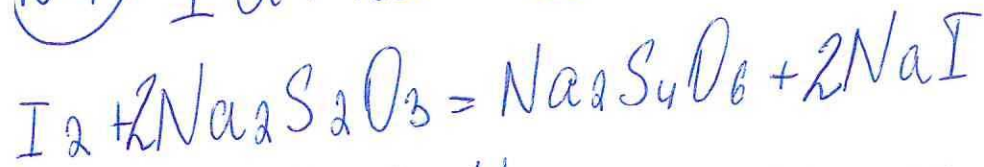
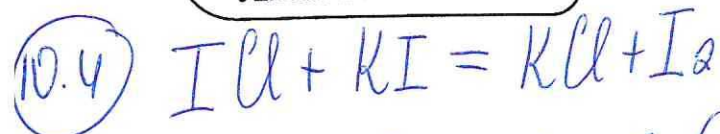


1)  $pV = \nu RT \Rightarrow \nu(CO_2) = \frac{101,5 \cdot 27,8}{8,31 \cdot 283} = 1,2 \text{ моль}$   
 2)  $12x + 6y = 1,2 \text{ моль}$   
 3)  $\nu(\text{мол. смеси}) = \frac{1,124}{22,4} = 0,05 \text{ моль} \Rightarrow \nu(Cu_2O) = 0,05 \text{ моль}$   
 4)  $\begin{cases} x + y = 0,05 \\ 12x + 6y = 1,2 \end{cases} \Rightarrow \begin{cases} x = \\ y = \end{cases}$



СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ





1)  $\nu(I\text{Cl}) = 20 \cdot 0,1 = 2$  ммоль

2)  $\nu(Na_2S_2O_3) = 0,25 \cdot 8,8 = 2,2$  ммоль

3)  $\nu(Na_2S_2O_3) = 0,25 \cdot 16 = 4$  ммоль

4)  $\nu(I_2) = \frac{2,2}{2} = 1,1$  ммоль. (8,4 г)

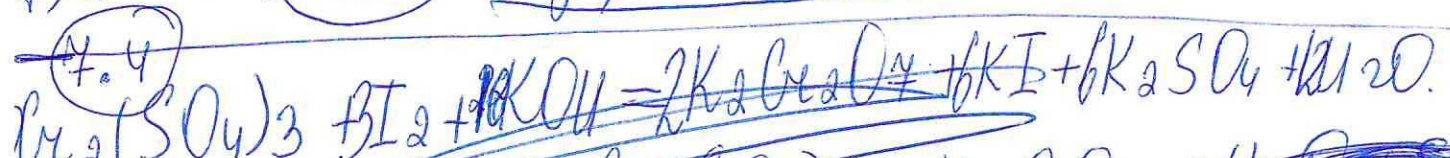
5)  $\nu(I_2) = 2 - 1,1 = 0,9$  ммоль

6)  $0,12 - 0,9$  ммоль  $\Rightarrow x = 900$  ммоль = 0,9 ммоль.

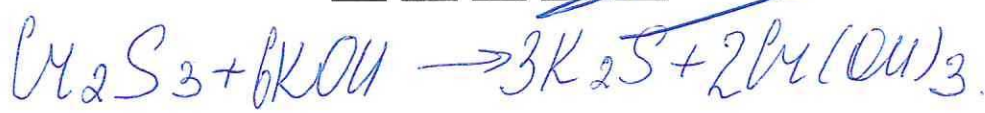
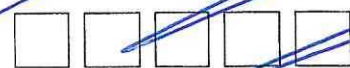
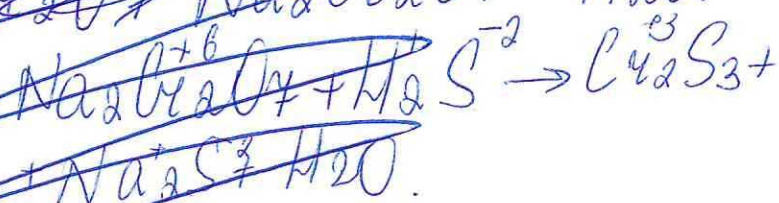
7)  $0,9 \cdot 127 = 114,3$  г. Не доброкочается вешое

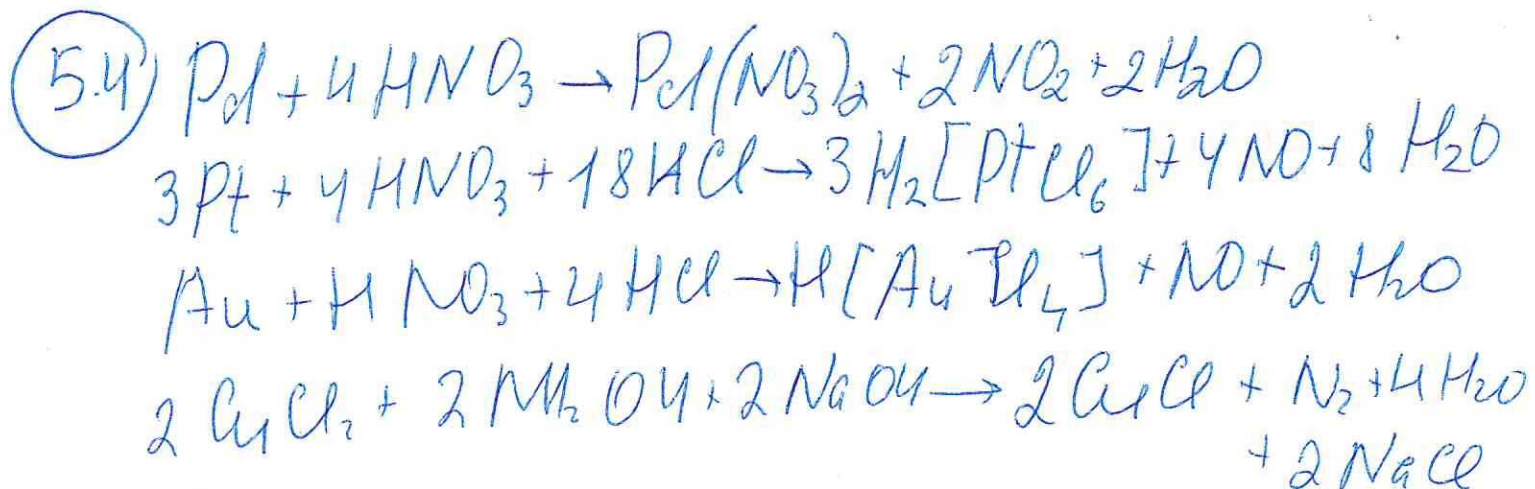
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СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ





$$V(\text{CuCl}_2) = 1,6$$

$$V(\text{N}_2) = 0,8$$

$$V_{\text{шара}} = \frac{4}{3}\pi r^3 = 4,443$$

$$m = V \cdot \rho = 57,9$$

$$V(\text{Pd}) = \frac{1}{2} V(\text{NO}_2) = 0,4$$

$$m(\text{Pd}) = 0,4 \cdot 106 = 42,4$$

$$m(\text{Pt/Au}) = 57,9 - 42,4 = 15,5$$

$$m(\text{Pd}) = 42,4$$

$$W = 73\%$$

$$m(\text{Pt}) = 9,8$$

$$W = 17\%$$

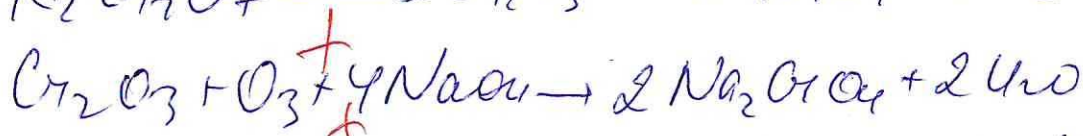
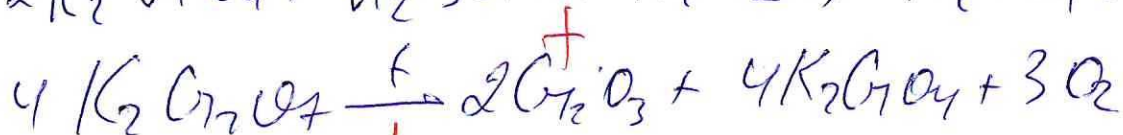
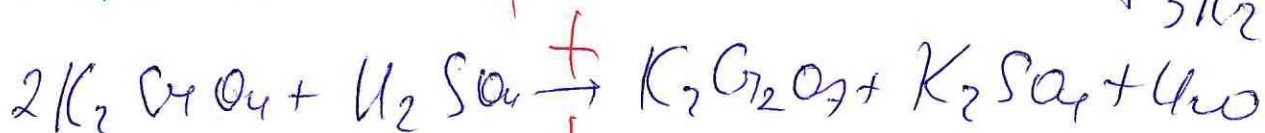
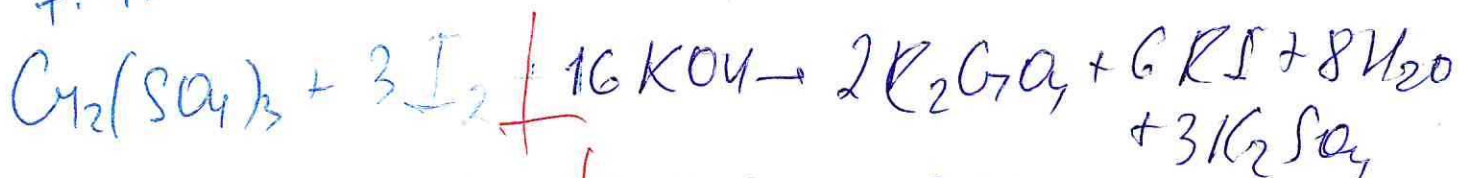
$$m(\text{Au}) = 5,7$$

$$W = 10\%$$

68



7.4.



+ 12e

