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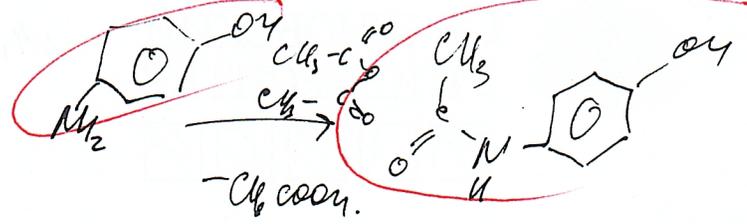
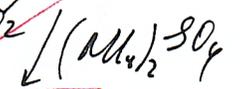
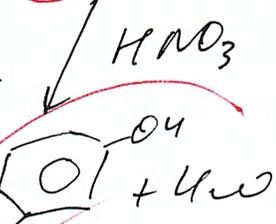
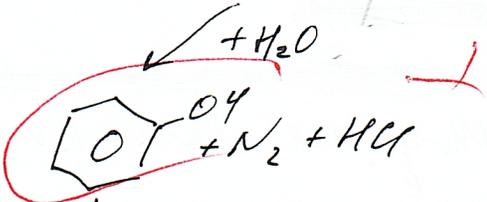
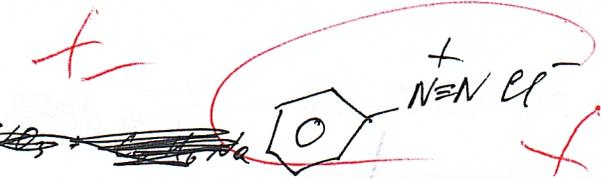
53  
/ /

6.2



60

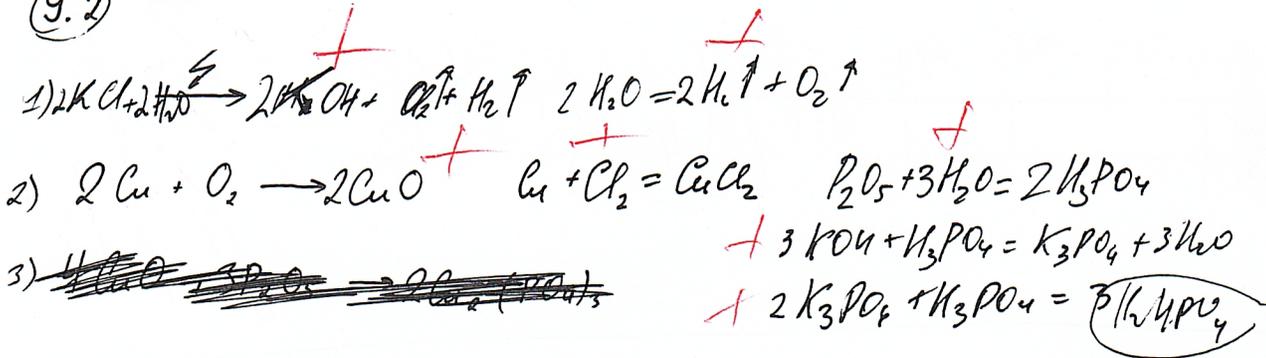
7.2



100



9.2



~~2)  $2Cu + O_2 \rightarrow 2CuO$~~

~~$4CuO + 3P_2O_5 \rightarrow 2Cu_2(P_2O_7)_3$~~

~~$m(Cu_2(P_2O_7)_3) = 2065,2$~~

$m_{p-pa} = 447 + 100 + 4,81 = 554,81$

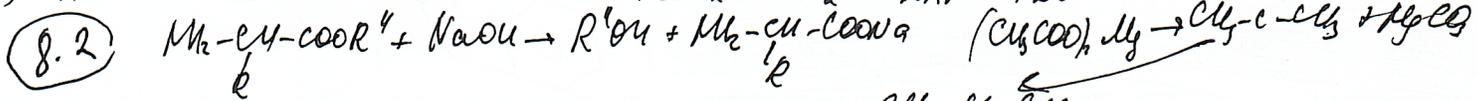
$w(Cu_2(P_2O_7)_3) = \frac{2065,2}{554,81} = 3,91\%$

75

5.2



2)  ~~$H_2SO_4 + Zn \rightarrow ZnSO_4 + H_2 \uparrow$~~



$n((C_2H_5COO)_2Mg) = \frac{36,8}{141} = 0,4$

$n(C_2H_5-C(=O)-C_2H_5) = 0,4 \Rightarrow n(-COOH) = 0,4$

$n(HCl) = \frac{21,9}{36,5} = 0,6$

$M_2-CH-COOC_2H_5$

$M = 231$      $n = \frac{1}{2}(-COOH) = 0,2$

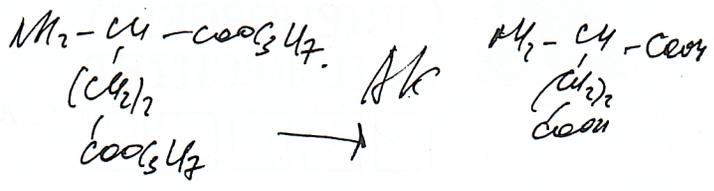
$M(x) = 28$      $(-C_2H_5)_2$

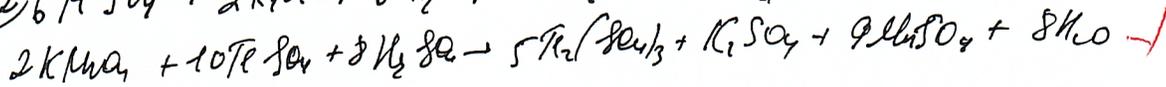
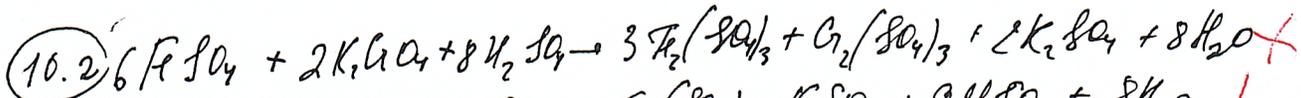
$M_2-CH-COOC_2H_5$      $M_2-CH-COONa$



СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ

11 X 0 0 4





$n(KMnO_4) = cV = 0,0025$

$n(FeSO_4) = \frac{10}{2} n = 0,0125$

$n(Fe_2SO_4) = cV = 0,45$

$n(K_2CrO_4) = 0,0025$

$m(K_2CrO_4) = 0,0025 \cdot 194 = 0,485$

12  
6250мг

