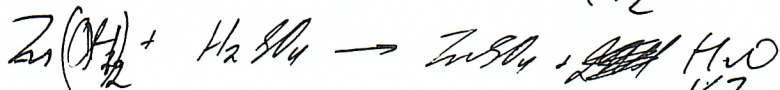
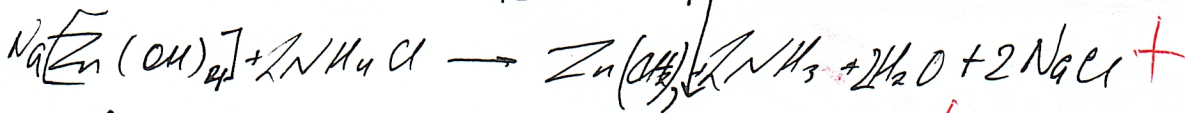
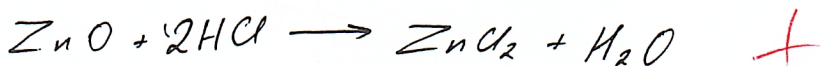
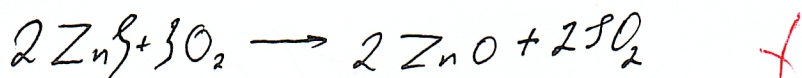


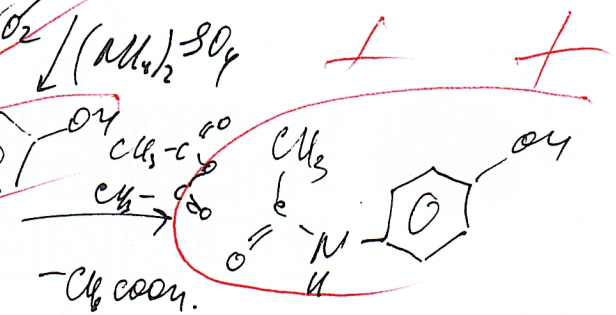
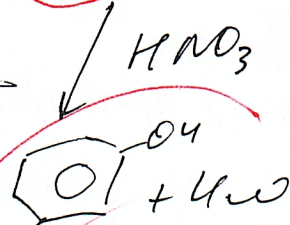
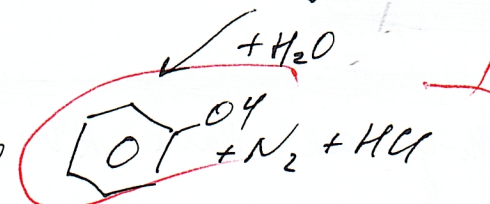
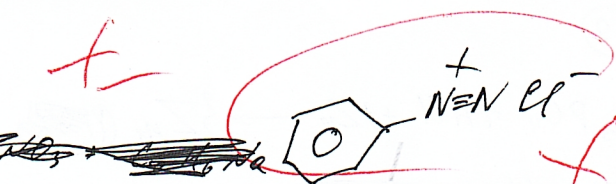
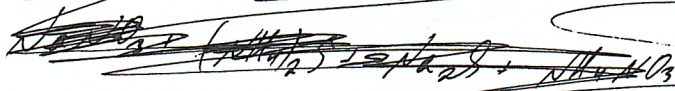
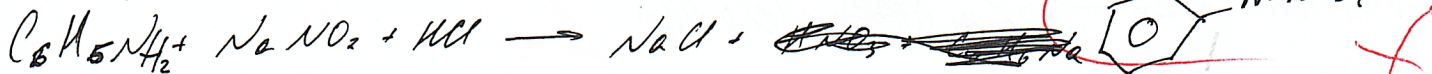
1	2	3	4	5	6	7	8	9	10
1				1	12	10	10	7	12

53 / 100

6.2



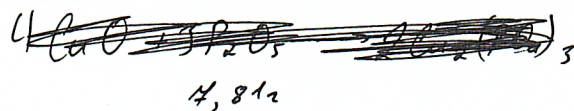
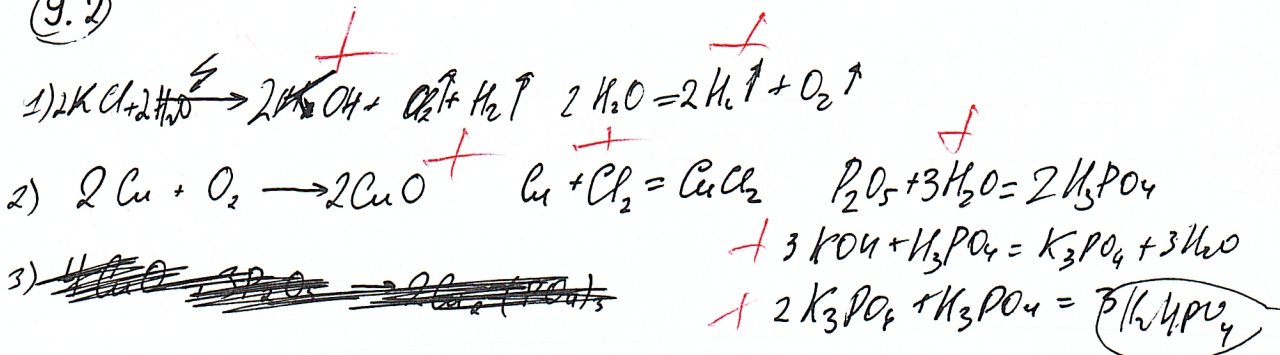
7.2



СЕЧЕНОВСКИЙ
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11X004

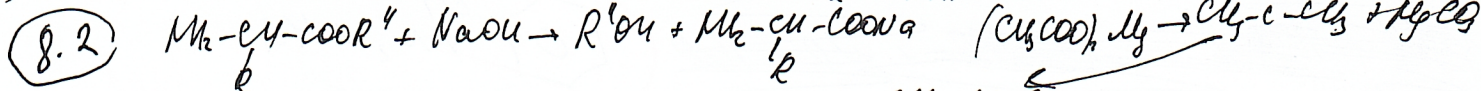
9.2



$m_{p-pa} = 447 + 100 + 4,81 = 551,81$

$w(Cu_3(PO_4)_2) = \frac{206,2}{551,81} = 3,74\%$

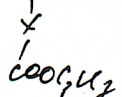
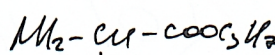
5.2



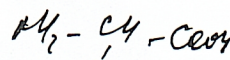
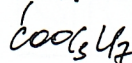
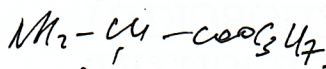
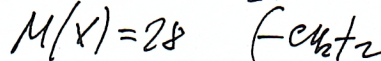
$n((CH_3COO)_2Mg) = \frac{36,8}{144} = 0,25$

$n(CH_3-CH_2-CH_3) = 0,4 \Rightarrow n(-COOH) = 0,4$

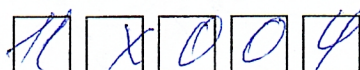
$n(HCl) = \frac{21,9}{36,5} = 0,6$

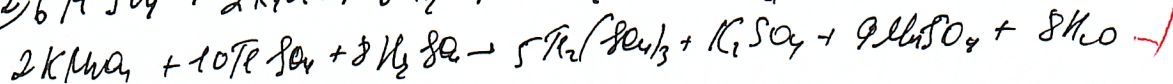
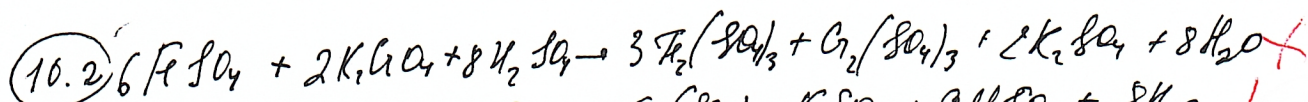


$M = 231$ $n = \frac{1}{2}(-COOH) = 0,2$



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$n(\text{KMnO}_4) = cV = 0,0025$

$n(\text{FeSO}_4) = \frac{10}{2} n = 0,0125$

$n(\text{FeSO}_4) = cV = 0,45$

$n(\text{K}_2\text{Cr}_2\text{O}_7) = 0,0025$

$n(\text{K}_2\text{Cr}_2\text{O}_7) = 0,0025 \cdot 104 = 0,26$

12
6250ms



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