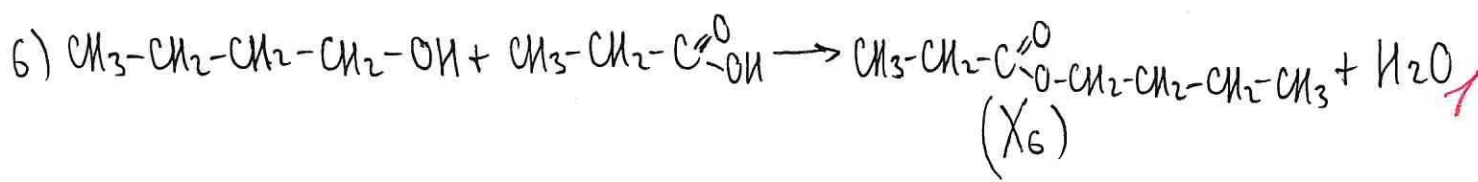
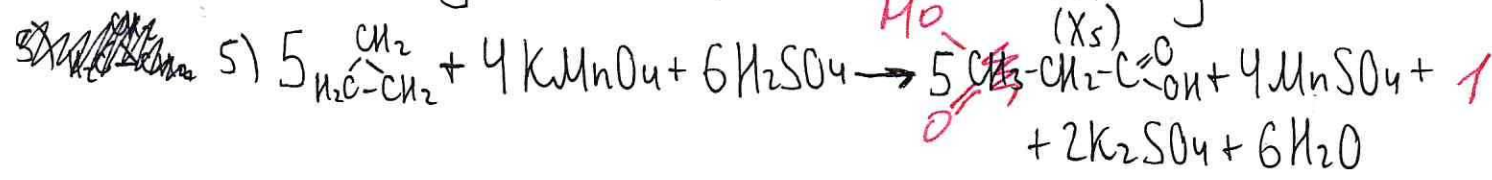
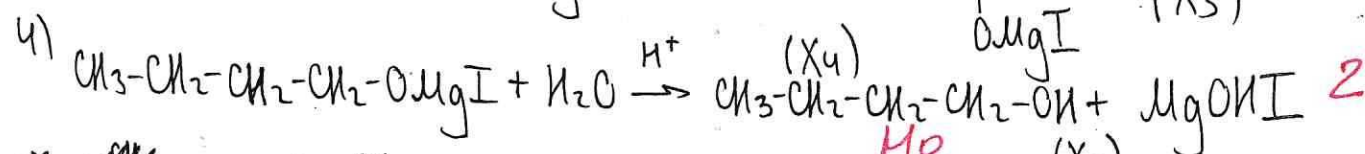
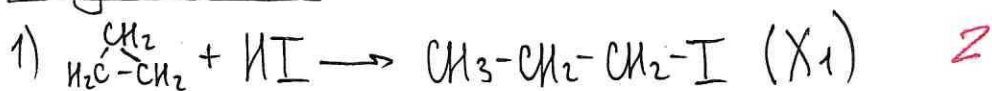


1	2	3	4	5	6	7	8	9	10
0	0	0	0	2	2	10	12	4	12

≤ 42

Задача 7.4



Задача 8.4.

$$M(NF_3) = 20 \frac{г}{моль} \cdot 3,55 = 71 \frac{г}{моль} \quad 2$$

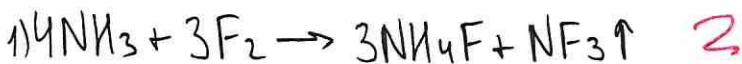
$$M(смеси (NH_3 \text{ и } F_2)) = 1,161 \frac{г}{л} \cdot 22,4 \frac{л}{моль} = 26,0064 \frac{г}{моль} \quad 2$$

$NF_3$

$$14 + 19x = 71$$

$$19x = 57$$

$$x = 3 \Rightarrow NF_3 = NF_3$$



По 1-ой р-ции:  $\frac{\nu(NH_3)}{\nu(F_2)} = \frac{4}{3}$  Пусть  $\nu(NH_3) = x$ , а  $\nu(F_2) = y$ , тогда  $y = \frac{x \cdot 3}{4}$   
(и по  $M(смеси)$ ):  $\frac{\nu(NF_3)}{\nu(F_2)} = \frac{y}{3} = \frac{3x}{3 \cdot 4} = \frac{x}{4}$

$$m(смеси) = \frac{58 \cdot 3x}{4} + \frac{85x}{4} = 64,75x \quad 2$$

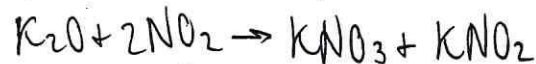
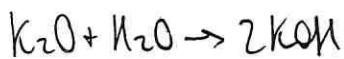
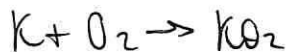
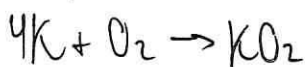
$$\omega(KNO_2) = \frac{21,25x}{64,75x} \cdot 100\% = 32,82\% \quad 2$$

$$\omega(KF) = 100\% - 32,82\% = 67,18\% \quad 2$$

12

Задача 9.4.

$$\nu(K) = \frac{17,55г}{39 \frac{г}{моль}} = 0,45 \text{ моль} \quad 2$$



4

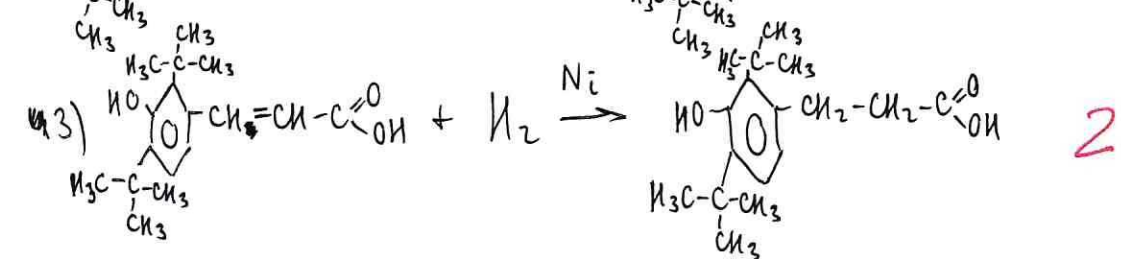
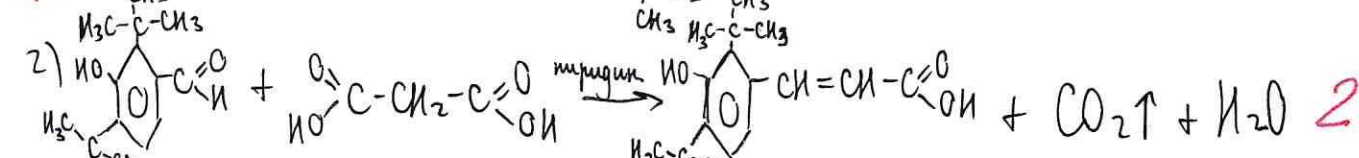
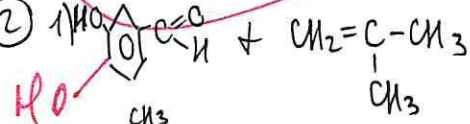
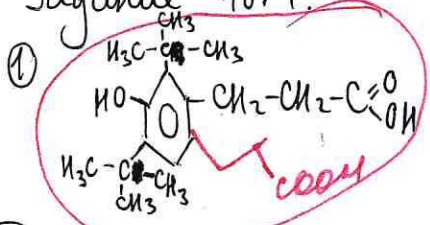


Задача 2.4.

$M(C_{13}H_{18}O_2N_2Cl_2) = 305 \frac{г}{моль}$   
 мелфалана

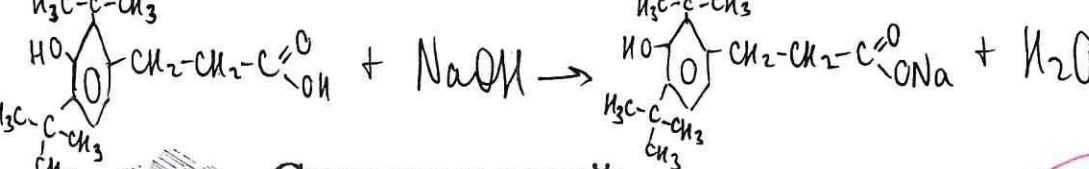
1 | 2 | 3 | 4 | 5 | 6 | 7 | 8 | 9 | 10  
 —————  
 | 12 | 4 | 12

Задача 10.4.



$V(C_7H_{12}O_3) = \frac{0,4}{278} = 0,0014 \text{ моль}$

$V(NaOH) = 2,8 \cdot 10^{-3} \cdot 0,1 = 0,28 \cdot 10^{-3} \text{ моль}$



12



Задача 5.4.

$$m(\text{Ca(OH)}_2) = 0,086 \text{ г}$$

$$\omega(\text{Ca(OH)}_2) = 0,086\%$$

$$\nu(\text{Ca(OH)}_2) = \frac{0,086 \text{ г}}{74 \frac{\text{г}}{\text{моль}}} = 0,0012 \text{ моль} \Rightarrow$$

$$\nu(\text{Ca}) = 0,0012 \text{ моль}$$

$$\nu(\text{OH}) = 0,0024 \text{ моль}$$

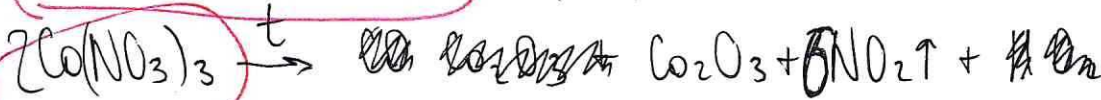
$$V_{\text{р-ра}} = 100 \text{ г}$$

$$K_p = [\text{Ca}^{2+}]^1 \cdot [\text{OH}^-]^2 = \frac{0,0012}{100} \cdot \left(\frac{0,0024}{100}\right)^2 = 1,2 \cdot 10^{-5} \cdot 5,76 \cdot 10^{-10} = 6,912 \cdot 10^{-15}$$

Задача 6.4.

$$M([\text{X}(\text{CO})_4]) = \frac{12 \cdot 4}{0,2812} = 170,67 \approx 171 \left(\frac{\text{г}}{\text{моль}}\right)$$

$$M(\text{X}) = 171 - 112 = 59 \left(\frac{\text{г}}{\text{моль}}\right) \Rightarrow \text{X} - \text{это Co}$$



NI!

