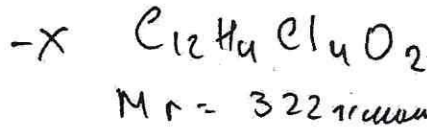
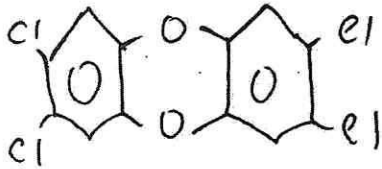
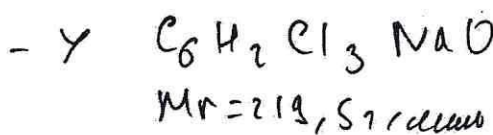
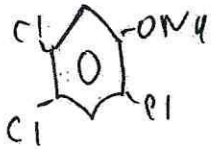


у 1



1/2/3/4/5/6/7/8/9/10  
 6/8/10/1-1-12/12/5-

Σ 43/5



$n(C) = 12x + 6y$

$12x + 6y = 2,7273(4x + 3y)$

$n(Cl) = 4x + 3y$

$12x + 6y = 10,9x + 8,2y$

$n(O) = 2,7273 \cdot n(Cl)$

$1,1x = 2,2y$

$x = 2y$

$m(C_{12}H_4Cl_4O_2) = 322x = 322 \cdot 2y = 644y$

$m(C_6H_2Cl_3ONa) = 219,5y$

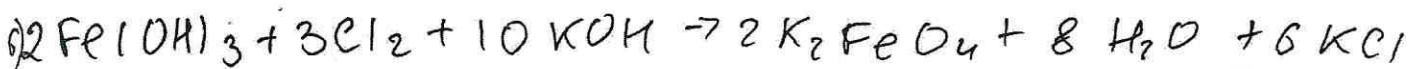
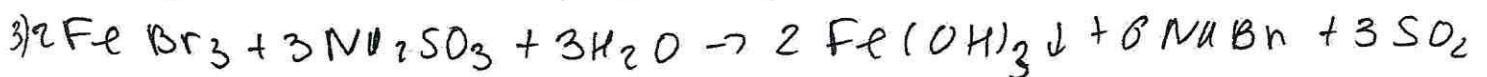
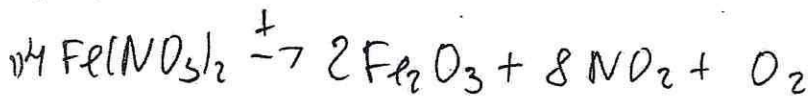
$w(C_{12}H_4Cl_4O_2) = \frac{644y}{644y + 219,5y} \cdot 100 = \frac{644}{863,5} \cdot 100 =$

$= 74,58\% \approx 74,6\%$

Σ 65

$w(C_6H_2Cl_3ONa) = 100 - 74,58 = 25,42\% \approx 25,4\%$

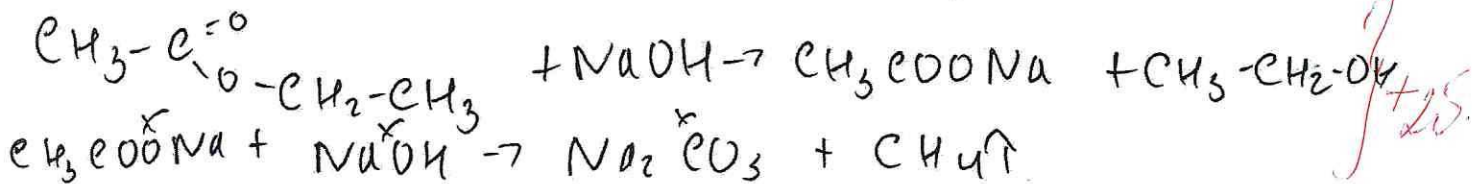
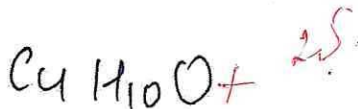
у 7



Σ 125

w2  
 $n(\text{CO}_2) = \frac{8,96 \text{ моль}}{22,4 \text{ моль}} = 0,4 \text{ моль}$     2)  $m(\text{C}) = 0,4 \cdot 12 = 4,8 \text{ г}$   
 3)  $n(\text{H}_2\text{O}) = \frac{a}{18} = 0,5 \text{ моль} \Rightarrow n(\text{H}) = 1 \text{ моль}$     4)  $m(\text{H}_2) = 1 \text{ моль} \cdot 2 = 2 \text{ г}$   
 5)  $m(\text{O}) = 7,4 - 4,8 - 2 = 0,6 \text{ г} \Rightarrow n(\text{O}) = \frac{0,6}{16} = 0,0375 \text{ моль}$

$n(\text{C}) : n(\text{H}) : n(\text{O})$   
 0,4    1    0,1  
 4    10    1



$n(\text{NaOH}) = 2,75 \cdot 0,091 = 0,25 \text{ моль}$   
 $m(\text{Na}_2\text{CO}_3) = 106 \text{ г/моль}$

$w(\text{Na}) \text{ в } \text{Na}_2\text{CO}_3 = \frac{46}{106} \cdot 100 = 43,4\% \Rightarrow \text{Na остаток}$

$\text{Na}_2\text{CO}_3 \left\{ \begin{array}{l} 23(2x+y) \\ 106x+40y \end{array} \right. = 0,45$   
*m сух. остатка = 26,00*

$m \text{ сух. остатка} = \frac{0,25 \cdot 23}{0,45} = 12,78 \text{ г} = 12,8 \text{ г}$

$2x + y = 0,25$   
 $106x + 40y = 12,8$

$\Rightarrow 106x + 40(0,25 - 2x) = 12,8 \Rightarrow x = 0,108$   
 $y = 0,035$

$n(\text{C}_4\text{H}_8\text{O}_2) = 0,75 - 0,108 = 0,642 \text{ моль}$

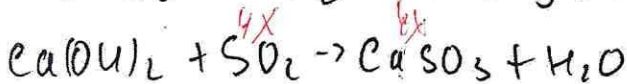
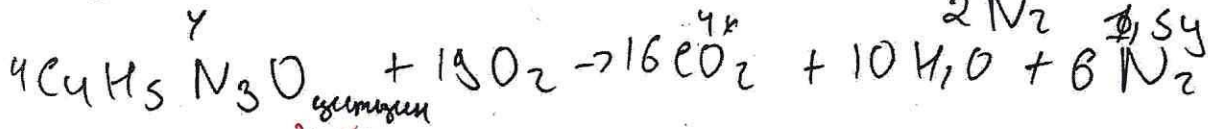
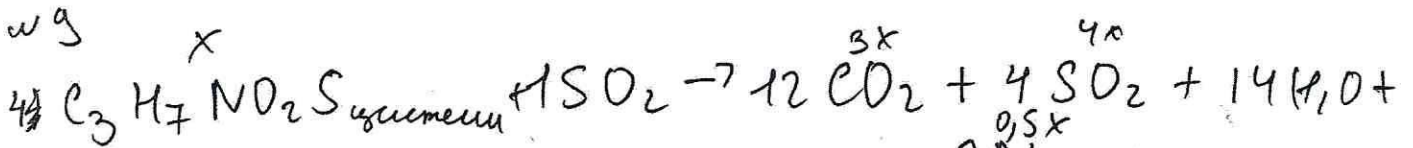
$m(\text{C}_4\text{H}_8\text{O}_2) = 0,642 \cdot 88 = 56,5 \text{ г}$

$m(\text{C}_4\text{H}_{10}\text{O}) = 15 - 56,5 = 2,5 \text{ г} \Rightarrow w(\text{C}_4\text{H}_{10}\text{O}) = \frac{2,5}{15} \cdot 100 = 16,7\%$



*x203*

*285*



$$0,5x = 1,5y$$

$$x = 3y$$

$$M(\text{C}_4\text{H}_5\text{N}_3\text{O}) = 111 \text{ г/моль}$$

$$M(\text{C}_3\text{H}_7\text{NO}_2\text{S}) = 121 \text{ г/моль}$$

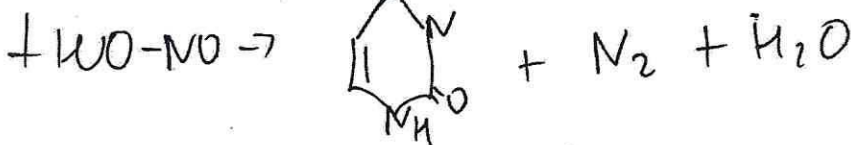
$$m(\text{CaCO}_3)_{\text{цистеин}} = (3x) \cdot 100 = 300x \Rightarrow \text{математика}$$

$$m(\text{CaSO}_3) = 4x \cdot 120 = 480x$$

$$m_{\text{осадка}} = 300x + 480x = 780x = 72940 \text{ г}$$

$$m(\text{CaCO}_3) = 4y \cdot 100 = 400y$$

$$\frac{m(\text{цитозин})}{m(\text{цистеин})} = \frac{2840 \text{ г}}{400 \text{ г}} = 7,1 \Rightarrow 7,1:1$$



$$n(\text{C}_4\text{H}_5\text{N}_3\text{O})$$

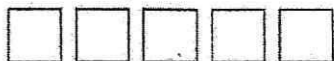
$$\frac{242}{111} = 2,18 \text{ моль}$$

$$\Rightarrow n_2 \Rightarrow V_{\text{N}_2} = 0,218 \cdot 22,4 = 4,88 \text{ л}$$

Σ 5

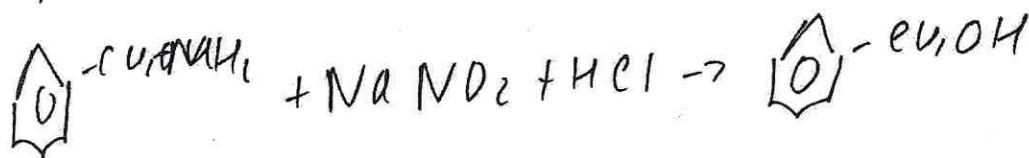
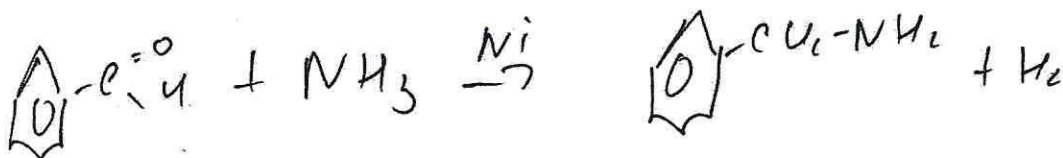
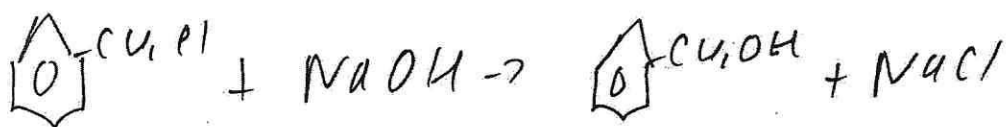
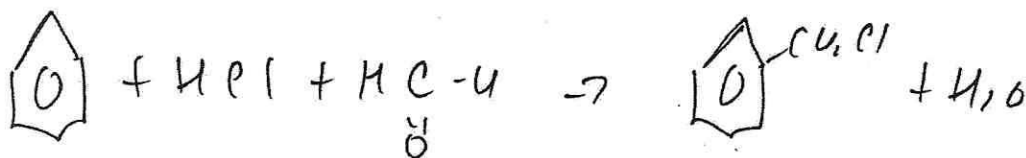


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УНИВЕРСИТЕТ

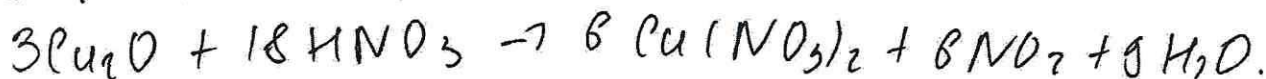
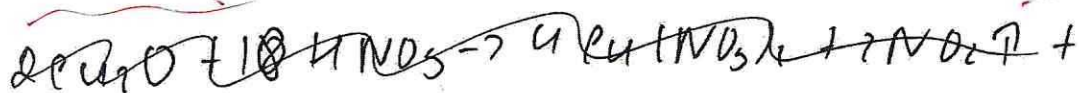
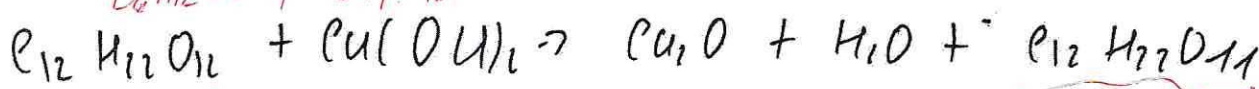


И.И. ДОВ

и 8



и 4  $\text{CaH}_2\text{O}_6 + \text{Cu}(\text{OH})_2 = ?$

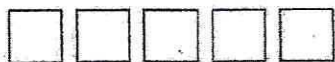


Σ 205

05



СЕЧЕНОВСКИЙ  
УНИВЕРСИТЕТ



11x 203