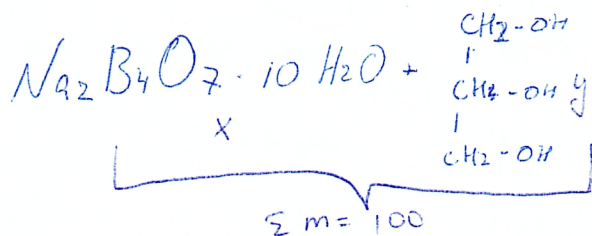


$$\begin{array}{r} 1 \ 7 \ 3 \ 4 \ 5 \ 6 \ 7 \ 8 \ 9 \ 10 \\ \hline 6 \ 6 \ 6 \ 6 \ 6 \ 0 \ 2 \ 10 \ 14 \end{array}$$

548 Жел.

2.3



1) По условию: $w\% \text{H} = \frac{m_{\text{b-l}}}{m_{\text{p-p}}} \cdot 100\% = 8\% \Rightarrow$

$$\begin{array}{l} \frac{28}{m_{\text{p-p}}} = \frac{382}{m_{\text{p-p}}} \\ 382x + 92y = 100 \\ 20x + 8y = 8 \\ m_{\text{b-l}} \end{array}$$

2) $w\% \text{B} = \frac{m_{\text{b-l}}}{m_{\text{p-p}}} \cdot 100\% = \frac{11.4}{100} \cdot 100\% \approx 11.4\%$

Ответ: 23%

3.3

$$w\% \text{смеси} = \frac{m_{\text{b-l}}}{m_{\text{p-p}}} \cdot 100\% = \frac{50.56}{135.7} \cdot 100\% = 37\%$$

$$\frac{50.56}{135.7} \cdot 100\% \approx 37\%$$



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N13



$$\sum n(\text{P}) = 2x + 7y = 0,7$$

$$\sum n(\text{H}) = 8x + 9y = 2,85$$

$$= 610 - 568 = 1172$$

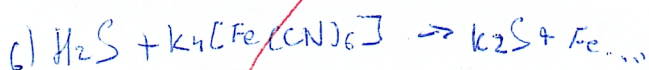
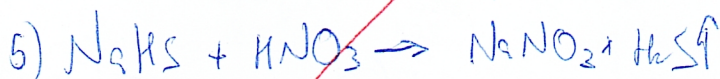
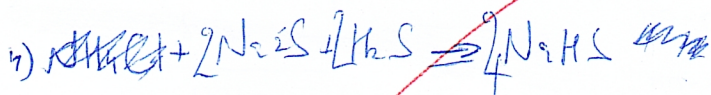
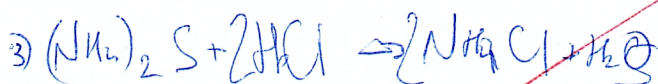
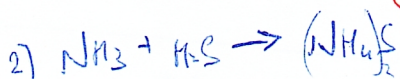
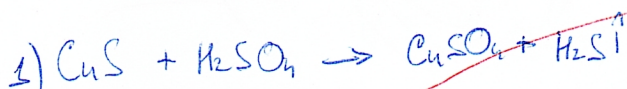
$$m(\text{MgZn}_2(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}) = 124,8$$

$$m(\text{Al}_3(\text{PO}_4)_2(\text{OH})_3 \cdot 3\text{H}_2\text{O}) = 18,8$$

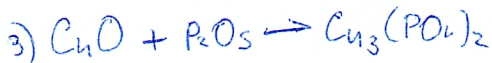
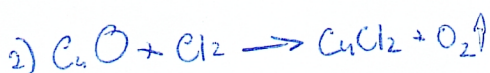
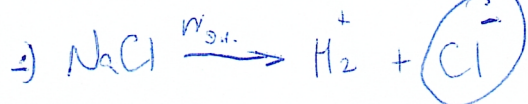
$$\sum m = 143,6$$

+6

N63



N93



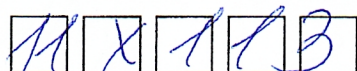
$$\omega_{\text{H}_2\text{O}} = \frac{m_{\text{H}_2\text{O}}}{m_{\text{проб}}} \cdot 100\% = 12,62\%$$

$$\mu\text{H}_2\text{O} = \frac{m_{\text{H}_2\text{O}}}{100} \cdot 100 = 12,63$$

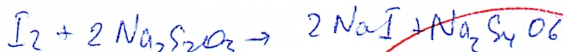
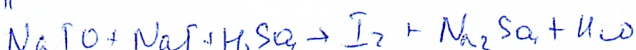
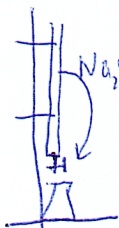
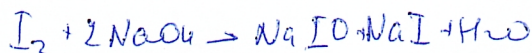
$$\text{Масса O} = 12,63\%$$



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N 10.3



+ 140

$$n(I_2) = \frac{1}{2} n(Na_2S_2O_3) = \frac{CV}{2} = 0,00075$$

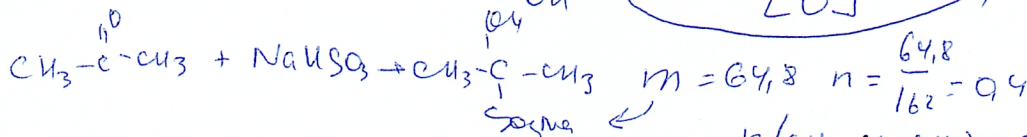
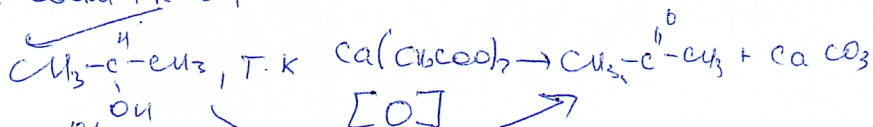
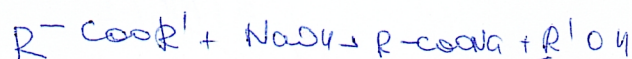
$$n(NaIO) = n(I_2) = 0,00075$$

$$n(H-C(=O)-OH) = 0,00075$$

$$m(H-C(=O)-OH) = 0,00075 \cdot 100 = 0,075 \text{ г}$$

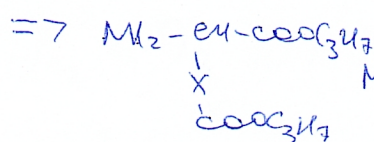
$$m = 1,125 \text{ г}$$

N 8.3



$$n(HBr) = \frac{48,6}{81} = 0,6$$

$$n(HBr) : n(-COOH) = 0,6 : 0,4 = 3 : 2$$

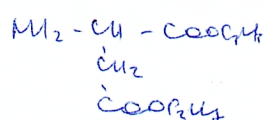


$$M = 231$$

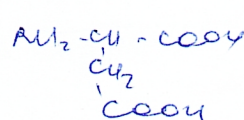
$$n(CH_3COOH) = \frac{1}{2} n(-COOH) = 0,2$$

$$M = 217$$

$$M(X) = \frac{CH_3}{14}$$



AK



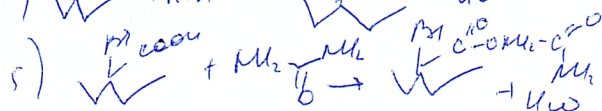
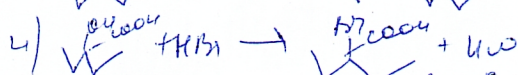
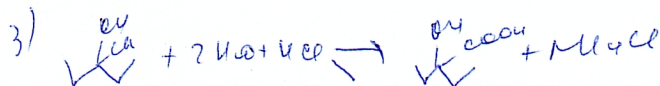
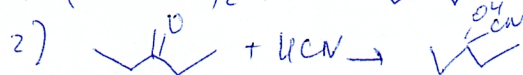
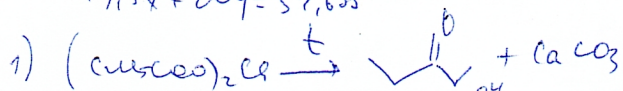
1025

N 7.3

$$x : y = 4 : 1$$

$$36,5x + 40y = 100$$

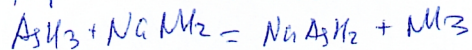
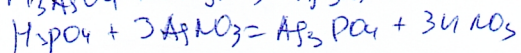
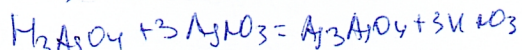
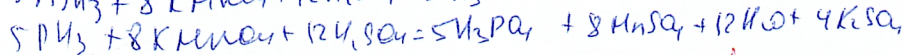
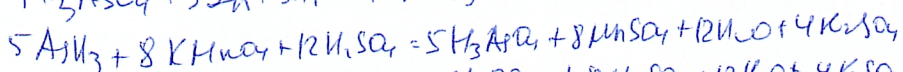
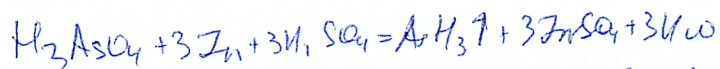
$$19,5x + 20y = 57,68$$



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N 5.3



$$M(\text{AsH}_3/\text{PH}_3) = \frac{\sum m}{\sum n} = 48,67$$

$$M = D \cdot M_{\text{H}_2} = 1,7382 \cdot 28 = 48,67$$

$$V = \frac{nRT}{p} - ?$$

$$n(\text{Na}_2\text{AsO}_4) = x \rightarrow n(\text{AsH}_3)/x$$

$$y = 2x$$

68



СЕЧЕНОВСКИЙ
УНИВЕРСИТЕТ

11113