54 / 68

1	2	3	4	5	6	7	8	9	10
6	6	8	/	/	12	12	10	/	/

1.1 $n(MgZn_2(PO_4)_2 \cdot 4H_2O) = x$ $n(Al_2(OH)_3PO_4) = y$

$n(O) = 12x$

$n(P) = 2x$

$n(O) = 7y$

$n(P) = y$

$17x + 7y = 1{,}27$

$2x + y = 0{,}21$

$x = 0{,}1$

$m(MgZn_2(PO_4)_2 \cdot 4H_2O) = 416 \cdot 0{,}1 = 41{,}6$

$m(Al_2(OH)_3PO_4) = 2$

$\Sigma m = 100$

+ 68

1.2 $n(Na_2B_4O_7 \cdot 10H_2O) = x$ $n(C_3H_8O_3) = y$

$382x + 92y = 100$

$20x + 8y = 8$

$m(C_3H_8O_3) = nM = 79{,}9$ $\omega = 79{,}9 \longrightarrow 80\%$

+ 68

4.1 $A + 6O + HNO_3 \longrightarrow NH_4NO_3 + CO_2 + NO_2 + H_2O$

$n(C) = n(CO_2) = PV/RT = 0{,}01$

$n(N) = n(NH_4NO_3) = m/M = 0{,}005$

$n(NO_2) = n(HNO_3) = 0{,}03$

$n(NH_3) = n(NH_4NO_3) = 0{,}005$

$\Sigma n = 0{,}035$ $0{,}035 + y = 0{,}04 - 0{,}04$

$n(H) = 0{,}013$ $m(O) = m_в - m(C) - m(H) - m(N) = 0{,}064$

$n(O) = 0{,}004$

$C_{10}H_{13}N_5O_4$

$\Sigma n\bar{e} = 6 \cdot 10 + 13 \cdot 1 + 5 \cdot 7 + 4 \cdot 8 = 140$

+ 8

6.1 $2ZnS + 3O_2 \rightarrow 2ZnO + 2SO_2$ +

$ZnO + 2NaOH + 2H_2O \rightarrow Na_2[Zn(OH)_4]$ +

$Na_2[Zn(OH)_4] + 2NH_4Cl \rightarrow Zn(OH)_2 + 2NaCl + 2NH_3 + 2H_2O$ +

$Zn(OH)_2 \rightarrow ZnO + H_2O$ +

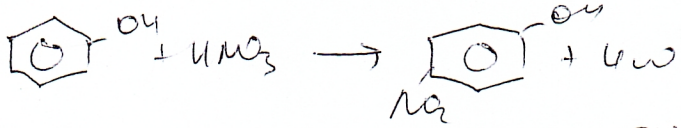
$ZnO + Co(NO_3)_2 \rightarrow CoZnO_2 + 2NO_2 + \frac{1}{2}O_2$ +

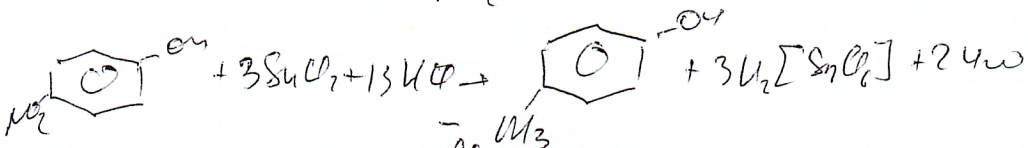
$CoZnO_2 + 6H_2SO_4 \rightarrow Co_2(SO_4)_3 + 2ZnSO_4 + SO_2 + 6H_2O$ +

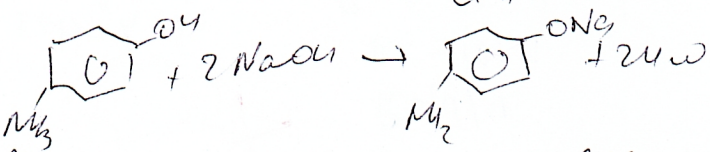
125

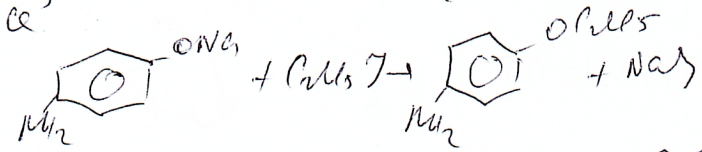
7.1 $13x + 14y = 100$ | $\times 6$

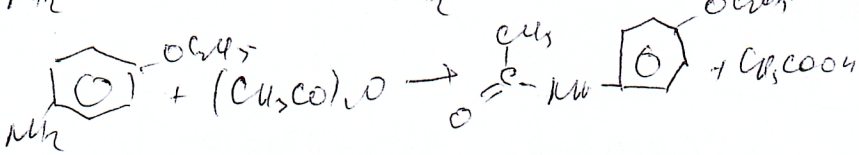
$7x + 8y = 532$ | $y = 11$



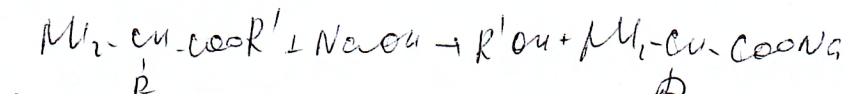


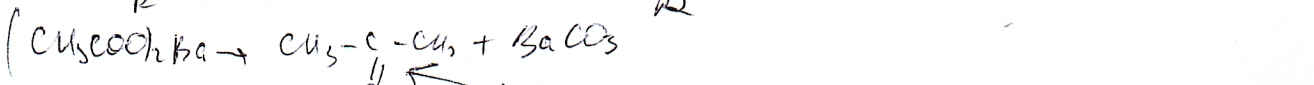






+125

8.1 



$\nu((CH_3COO)_2Ba) = \frac{m}{M} = 0,8 \rightarrow$  $\rightarrow m(CH_3-CH(OH)-CH_3) = 0,8 \rightarrow \nu(-COOC_3H_7) = 0,8$

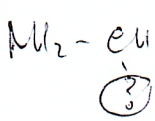
$\nu(HBr) = \frac{m}{M} = 1,2$

$\nu(COOC_3H_7) : \nu(HBr)$

$0,8 \quad 1,2$

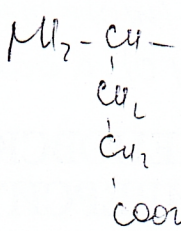
$2 : 3$

$= -COOH$?



$M = \frac{92,4}{0,4} = 231$

$M(?) = 28$



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