



$$y = z$$

$$\frac{1}{3}x = y = z$$

$$2y = y = z = \frac{1}{3}x$$

Ответ:

$$C = 12 \text{ ат}$$

$$H = 4 \text{ ат}$$

$$H = 4 \text{ ат}$$

$$O = 2 \text{ ат}$$



8

Задача 1.1

$$2y = \frac{1}{3}x$$

$$y = \frac{1}{6}x$$

$$y = 0.166x$$

$$\frac{1}{3}x + \frac{1}{3}x + x + 0.166x = 22$$

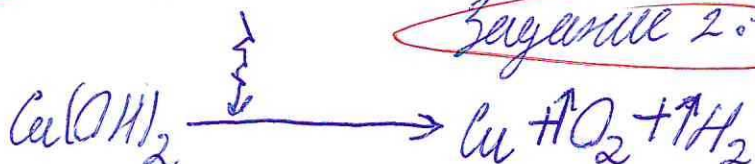
$$0.33x + 0.33x + x + 0.166x = 22$$

$$0.666x + x + 0.166x = 22$$

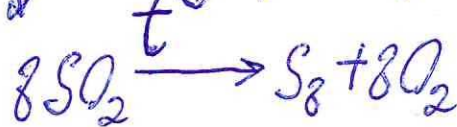
$$1.820x = 22$$

$$x = 12$$

Задача 2.1



↓ - микрограмм



Задача 3.1

$$\begin{array}{ccc|ccc|ccc} 1 & 2 & 3 & 4 & 5 & 6 & 7 & 8 & 9 & 10 \\ 8 & 4 & 4 & 2 & 10 & 2 & 5 & - & - & 0 \end{array}$$

Σ 37



~~Зачетная Ч.1~~

~~$M_{H_2SO_4} = 2 + 32 + 64 = 98$ г/моль~~

~~M_{KNO_3} г/моль~~

~~KBr - окислитель галогена превращаясь ок. в, реак. группа ок.~~

~~Значит мм ~~не~~ окисляется бром до Br₂ и выше~~

~~это может сделать HCl, H₂SO₄, HNO₃, H₃PO₄ и т.д.~~

~~$M_{HCl} = 36,5$ г/моль мм не нейтрал~~

~~Выбираем кислоты~~

~~F - фторид - изотоп воздуха 3~~

~~O - кислород - изотоп воздуха~~

~~Cl³⁵ - изотоп воздуха мм=2~~

~~Cl³⁸ - изотоп воздуха мм=2~~

~~H_2SO_3 H_2SO_4 H_2S~~

~~$T_2S = 32 + 3 \cdot 2 = 38$ г/моль~~

~~$P_2O_5 = 2 + 36 = 38$ г/моль~~

~~рек~~

~~HNO_3 $M = 63$ г/моль~~

~~H_2CO_3 $M = 63$ г/моль~~

~~O - кислород, изотоп мм=2~~



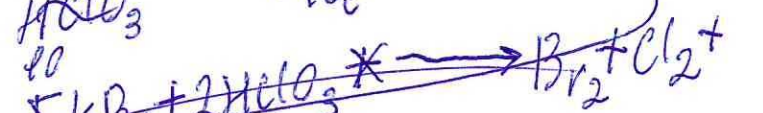
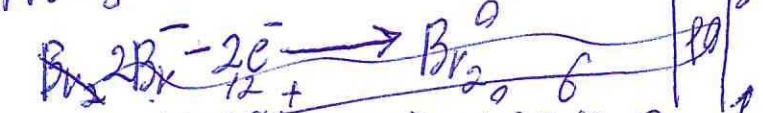
~~Выбираем кислоты~~

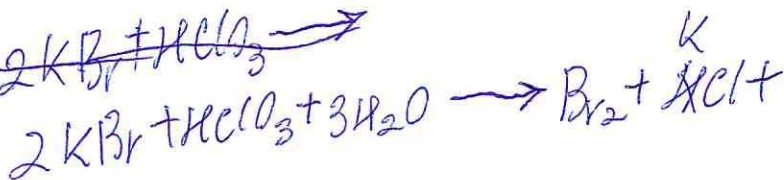
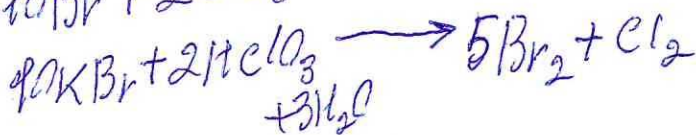
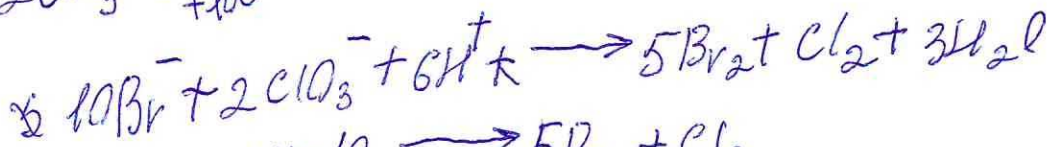
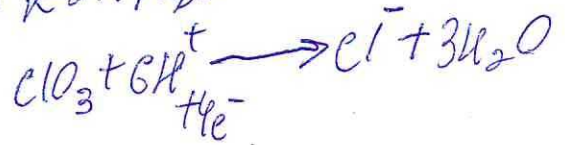
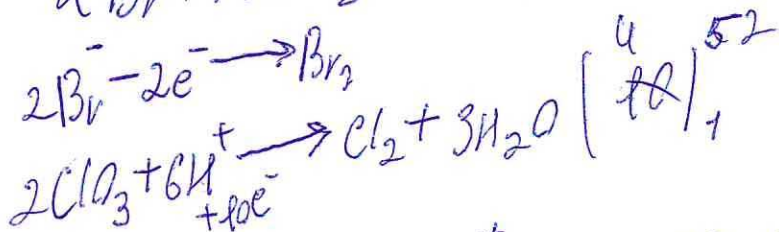
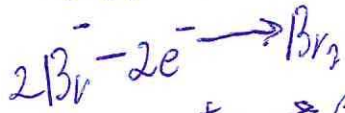
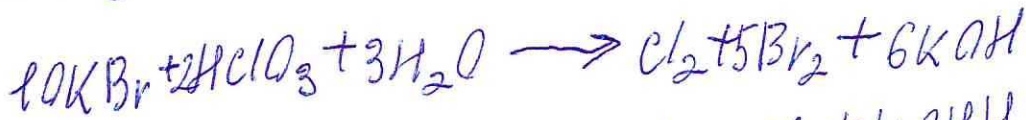
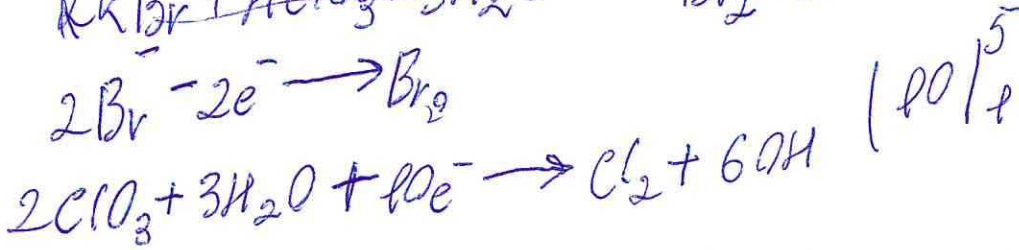
~~$HClO_3$ и $PTSO_3$~~

~~PT - изотоп воздуха~~

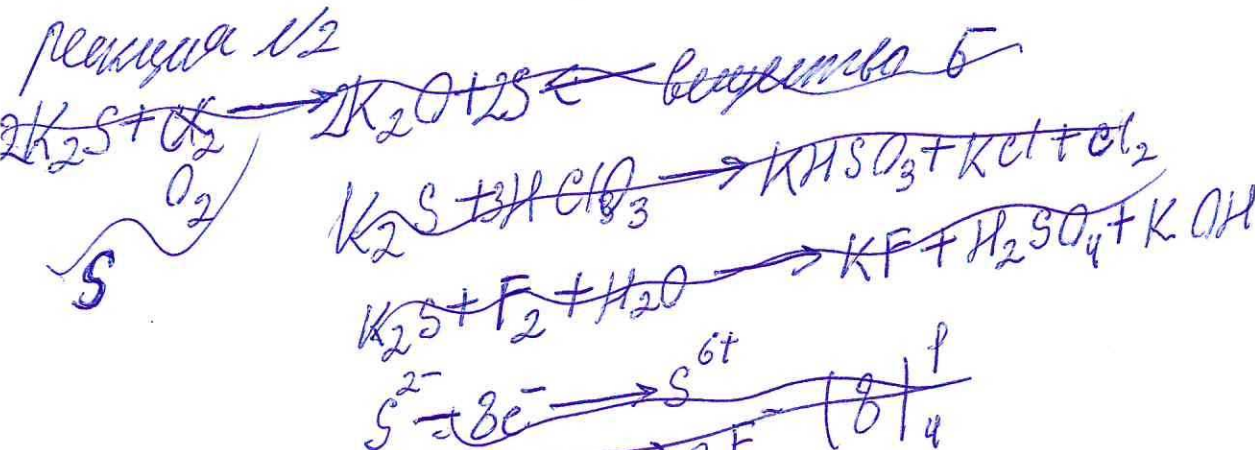
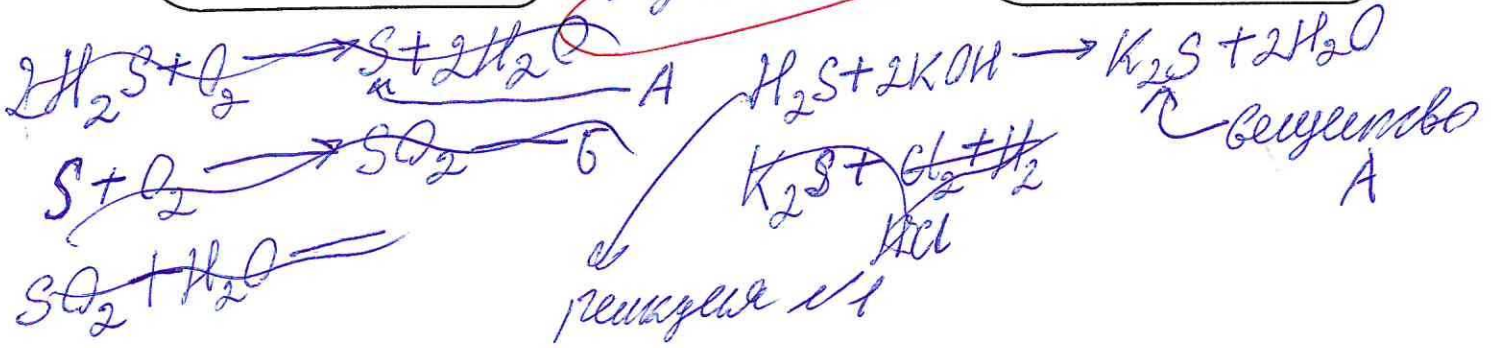
~~P M=2~~

~~T M=3~~

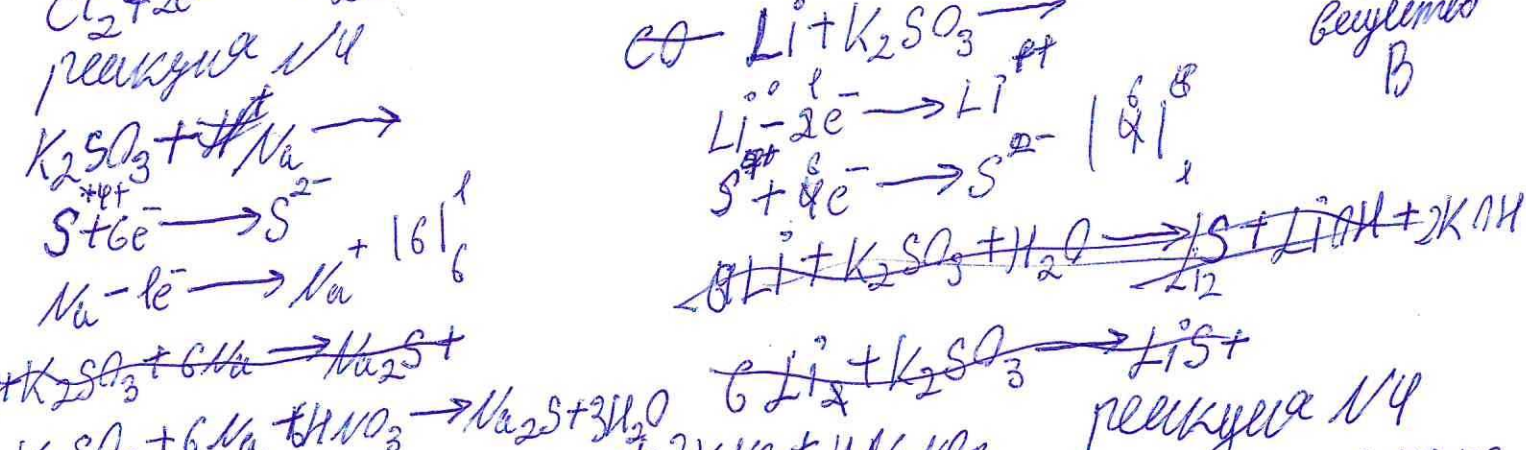
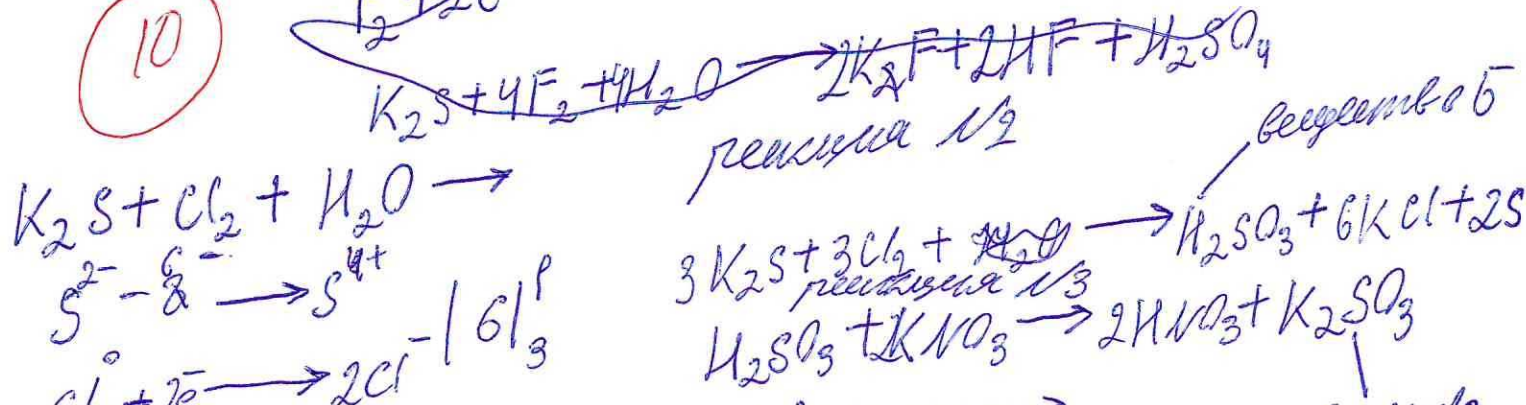




Задача 5.1

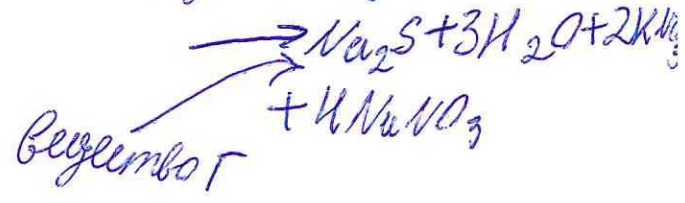


10

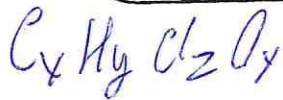


**СЕЧЕНОВСКИЙ
УНИВЕРСИТЕТ**

□ □ □ □ □



Задача 2.1



$$y=2$$

$$\frac{1}{3}x = y = z$$

$$2x = y = z = \frac{1}{3}x$$

$$2x = \frac{1}{3}y$$

$$y = \frac{1}{3}x/2$$

$$x = 0.166x$$

$$\frac{1}{3}x + \frac{1}{3}x + x + 0.166x = 22$$

$$0.33x + 0.33x + x + 0.166x = 22$$

$$0.666 + 0.166x = 22$$

$$1820x = 22$$

$$x = 12$$

Ответ:

$$C = 12 \text{ ат}$$

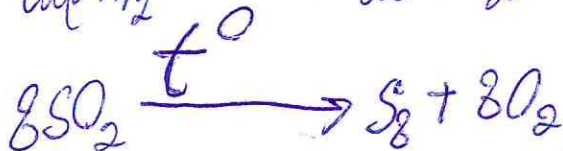
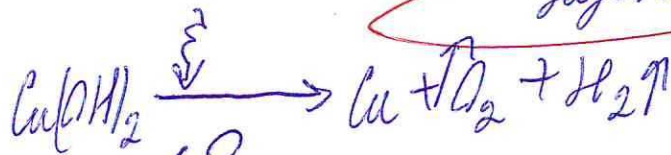
$$Cl = 4 \text{ ат}$$

$$H = 2 \text{ ат}$$

$$O = 2 \text{ ат}$$

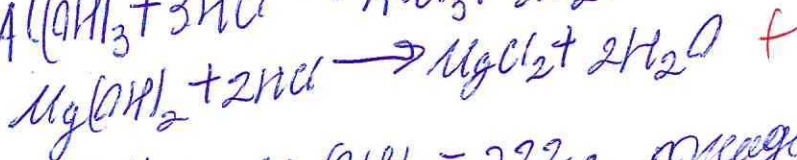
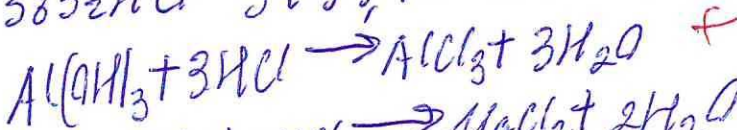


Задача 2.1



Задача 3.1

$$3832 \text{ HCl} \cdot 389,75 \cdot 0.12 = 3832 \text{ HCl} = 10.5 \text{ моль HCl}$$



$$m Al(OH)_3 + m Mg(OH)_2 = 282 \text{ г. масса балла}$$

$$10.5/3 = 3.5 \text{ моль } Al(OH)_3$$

$$10.5/2 = 5.25 \text{ моль}$$

$$3.5 \cdot 78 = 273 \text{ г. масса } Al(OH)_3$$

$$5.25 \cdot 58 = 304.5 \text{ г. масса } Mg(OH)_2$$

$$273x + 304.5y = 282$$

$$273x + 304.5 - 304.5x = 282$$

$$-31.5x = -225$$

$$x = 0.7$$



СЕЧЕНОВСКИЙ
УНИВЕРСИТЕТ



10.5 моль вес бисакей, 282 г. моль,
 $0.7 \cdot 273 = 191.1 \text{ г.}$ Al(OH)_3 $191/282 =$
 $0.3 \cdot 304.5 = 91.3$

Ответ: 0.677 и 0.323
 Al(OH)_3 Mg(OH)_2

Задача 8.1

$\frac{x \cdot 100}{\frac{x}{3} \cdot 16} = 1.5952$

$\frac{x+60}{3x+62} = 1.5952$

$\frac{3x+62}{x+60} = 1.5952$

$(3x+62) = (x+60) \cdot 1.5952$

$3x+62 = 1.5952x + 95.7$

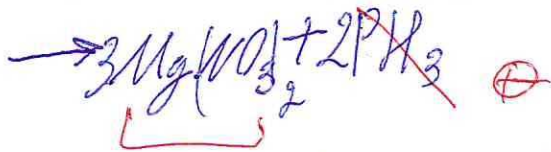
Ответ: $x = \text{Mg}$ 2.5

$3x - 1.6x = 95.7 - 62$

0.243

$1.4x = 33.7$

$x = 24 - \text{Mg}$

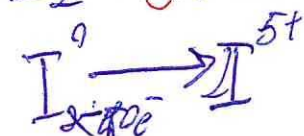
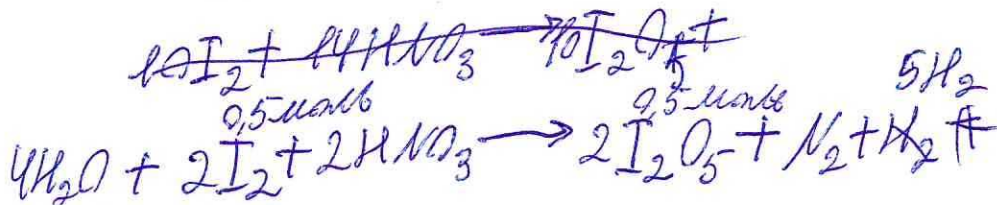
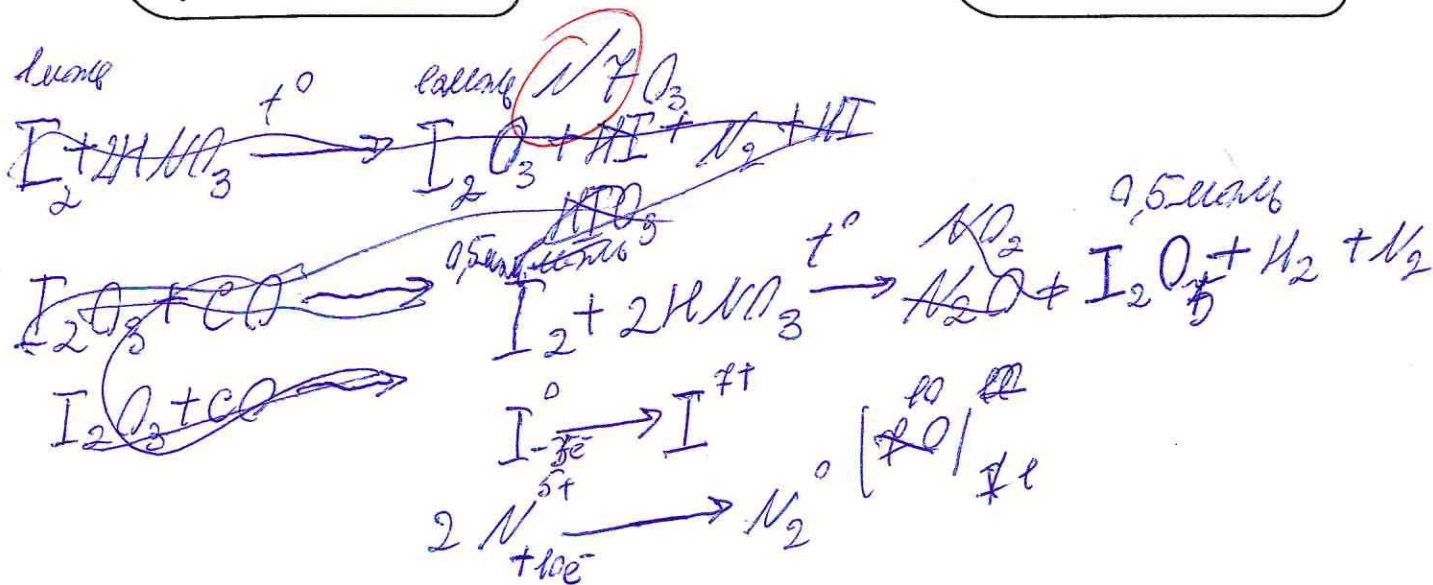


$26.8 / 24 + 24 + 31 \cdot 2 = 0.243 \text{ моль}$

$0.486 \cdot (31+3) = 16.5 \text{ г.}$

Ответ: масса газа
 будет 16.5 г.

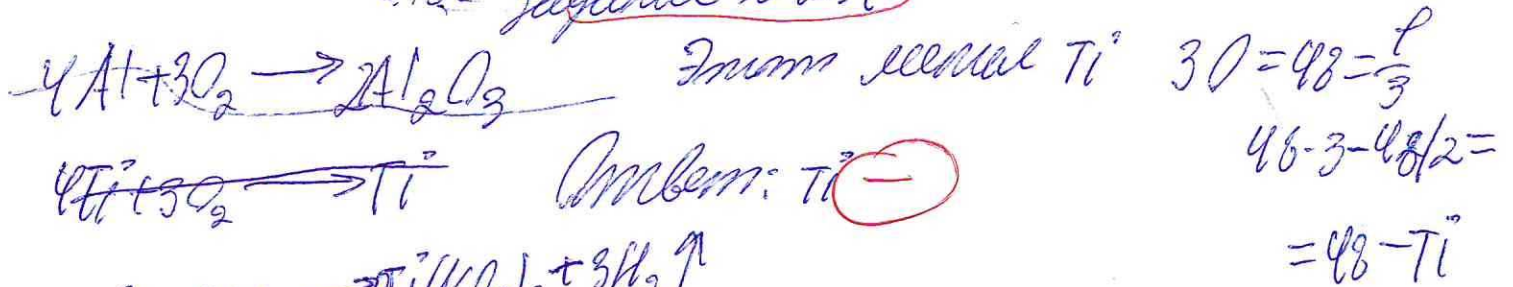




Ответ: KIO_3 3,52. & 1,12 KI
3,52,

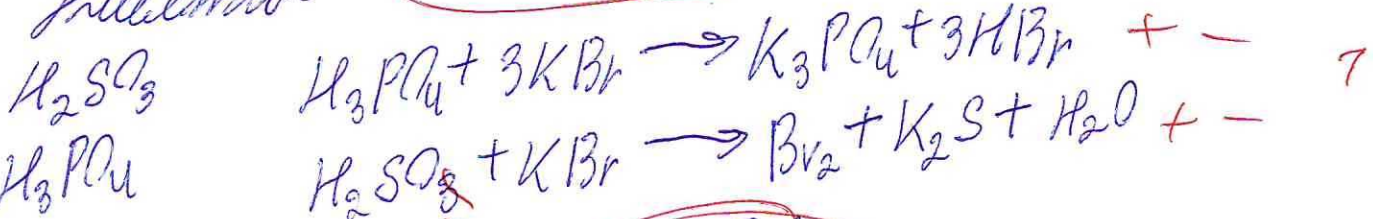


задание 11.1

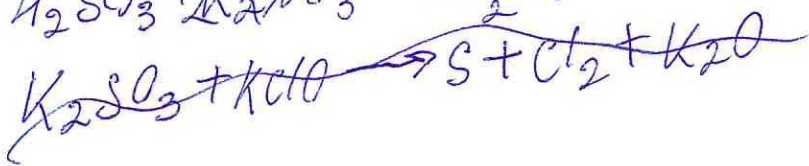
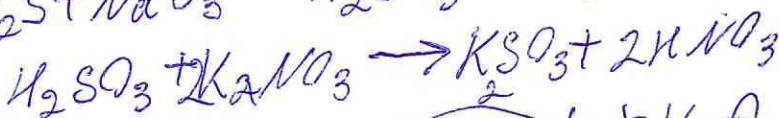
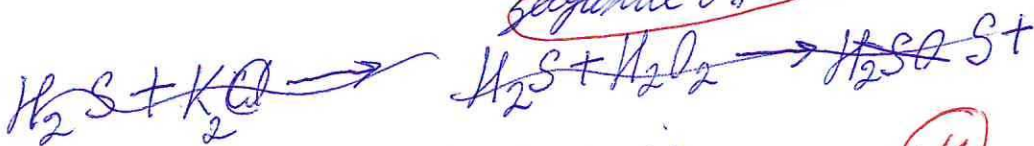


задание 4.1

Классификация



задание 5.1



(4)

