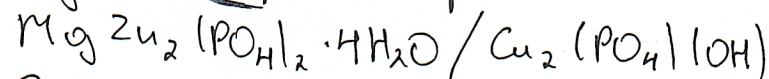


$$\begin{array}{cccccccccc}
 1 & 2 & 3 & 4 & 5 & 6 & 7 & 8 & 9 & 10 \\
 6 & 1 & 2 & 4 & 12 & 6 & 8 & 8 & &
 \end{array}$$

$\Sigma 47$

$$n(O) = \frac{1,5953 \cdot 10^{24}}{6,02 \cdot 10^{23}} = 2,65 \text{ моль} \quad n(P) = \frac{2,709 \cdot 10^{23}}{6,02 \cdot 10^{23}} = 0,45 \text{ моль} \quad \sim 1.2$$

Пусть $n(\text{серьсирновина}) = x \text{ моль}$, $n(\text{идемерина}) = y \text{ моль}$

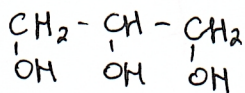
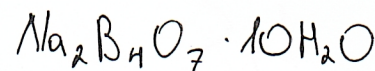


$$\begin{cases} 12x + 5y = 2,65 \\ 2x + y = 0,45 \end{cases} \Rightarrow x = 0,2 \text{ моль}, y = 0,05 \text{ моль}$$

$$M(\text{серьсирновина}) = 416 \text{ г/моль}$$

$$M(\text{идемерина}) = 240 \text{ г/моль}$$

$$m(\text{сб.р.}) = 0,2 \cdot 416 + 0,05 \cdot 240 = 95,2 \text{ г} \quad \text{Данное: } m(\text{сб.р.}) = 95,2 \text{ г.}$$



$$\begin{aligned}
 n(H) &= 28 \\
 M(H) &= 28 \text{ г/моль}
 \end{aligned}$$

~ 2.2

$$M(\text{b.ka}) = \frac{28}{0,08} = 350$$

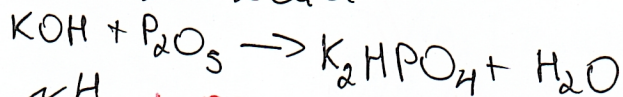
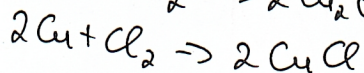
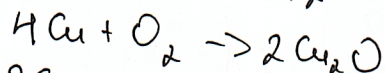
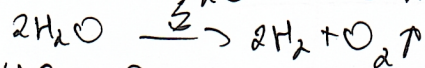
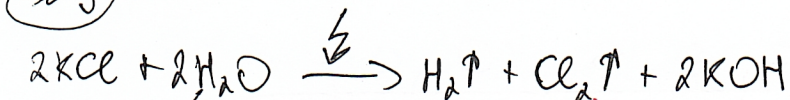
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~ 9



~ H H_3PO_4

$$n(KCl) = \frac{447,05}{74,5} = 0,3 \text{ моль}$$

$$n(Cu) = \frac{100}{64} = 1,5625 \text{ моль}$$

$$n(P_2O_5) = 0,055 \text{ моль}$$

$$PV = nRT$$

$$n(NO_2) = \frac{PV}{RT} = \frac{99,725 \cdot 1,293}{298 \cdot 8,314} = 0,052 \text{ моль}$$

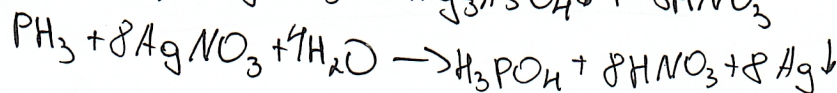
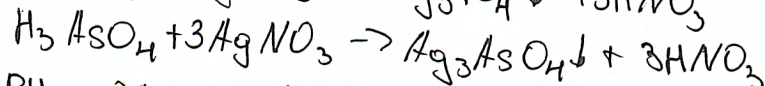
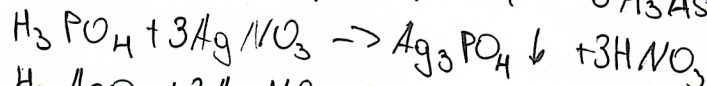
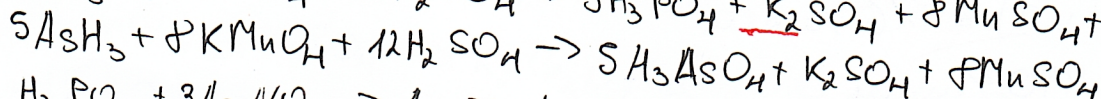
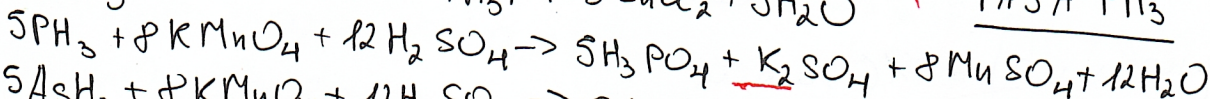
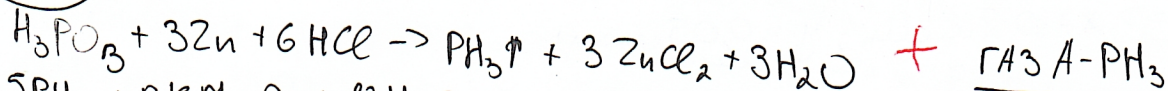
$$n(CO_2) = \frac{PV}{RT} = \frac{0,848 \cdot 99,725}{298 \cdot 8,314} = 0,014 \text{ моль}$$

$$n(CO_2) = 0,14 \text{ моль}$$

$$n(H_2O) = \frac{0,504}{18} = 0,028$$

$$n(NO_2) = 0,052 \text{ моль}$$

~ 5



ГАЗ А- PH_3

~~45~~ 0

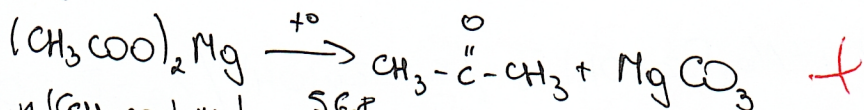
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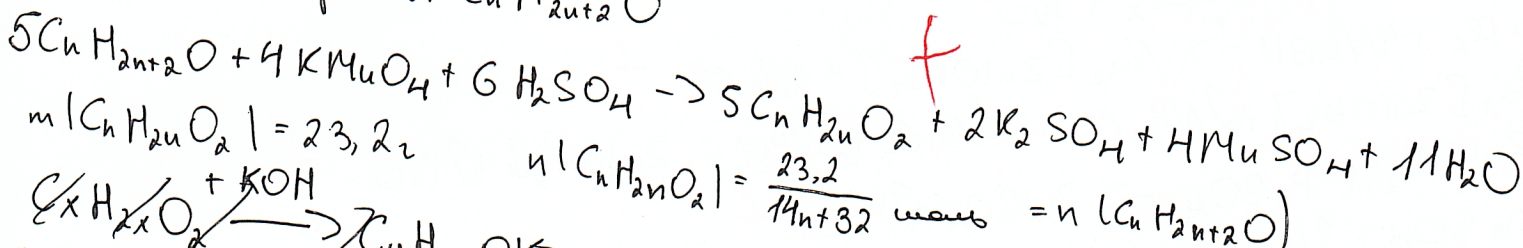
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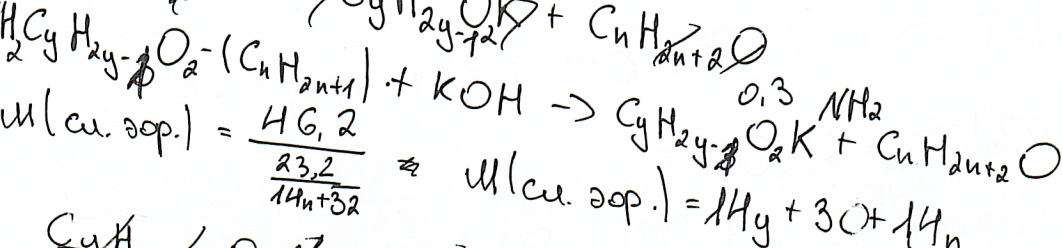
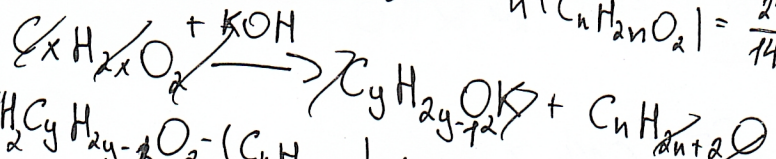
$$n(\text{CH}_3\text{COO})_2\text{Mg} = \frac{56,8}{142} = 0,4 \text{ моль} \Rightarrow n(\overset{\overset{\text{O}}{\parallel}}{\text{C}}) = 0,4 \text{ моль} \Rightarrow n(\overset{\overset{\text{O}}{\parallel}}{\text{C}}) = 0,4, \text{ SP} = 23,22.$$

Согл. ф-ла ацидов: $\text{C}_n\text{H}_{2n+2}\text{O}$

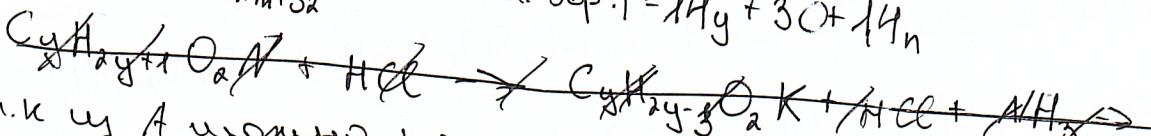


$$n(\text{C}_n\text{H}_{2n}\text{O}_2) = 23,22$$

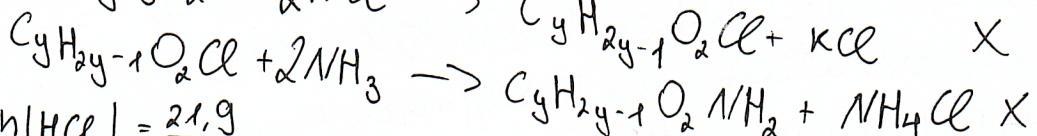
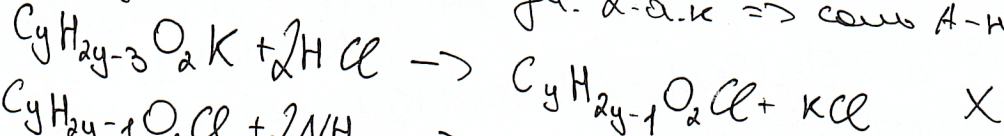
$$n(\text{C}_n\text{H}_{2n}\text{O}_2) = \frac{23,2}{14n+32} \text{ моль} = n(\text{C}_n\text{H}_{2n+2}\text{O})$$



$$M(\text{с. зор.}) = \frac{46,2}{\frac{23,2}{14n+32}} \approx M(\text{с. зор.}) = 14y + 30 + 14n$$

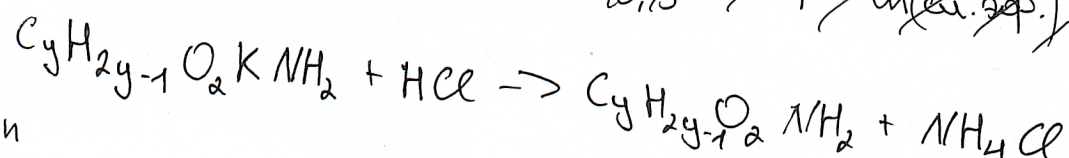


м.к. у А можно найти. д.а.к $\Rightarrow$ с.з. А-непределенная X



$$n(\text{HCl}) = \frac{21,9}{36,5} = 0,6 \text{ моль} \Rightarrow n(\text{A}) = 0,6 \text{ моль}$$

$$\frac{23,2}{14n+32} = 0,6 \Rightarrow n = 3, \quad \frac{46,2}{6,8} = 154 = M(\text{с. зор.})$$



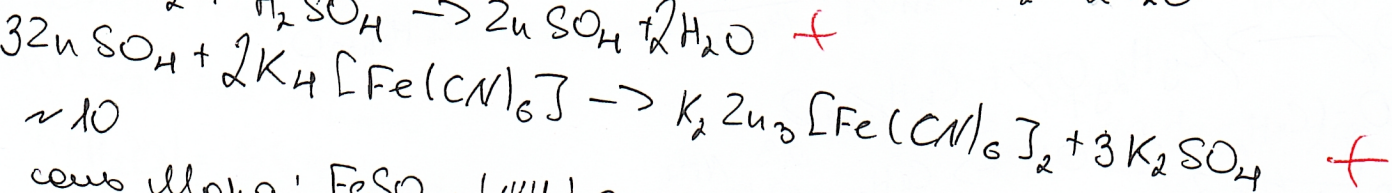
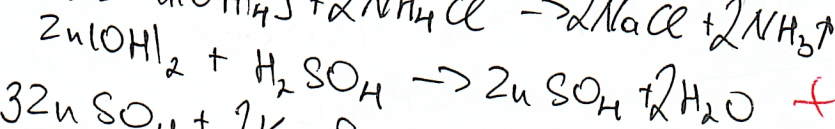
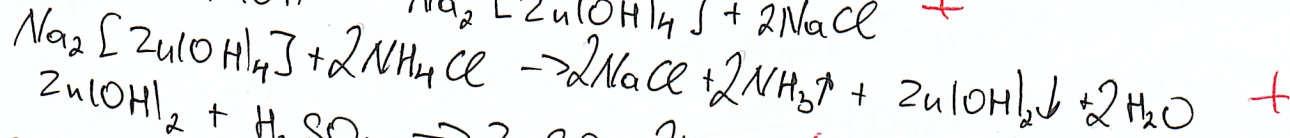
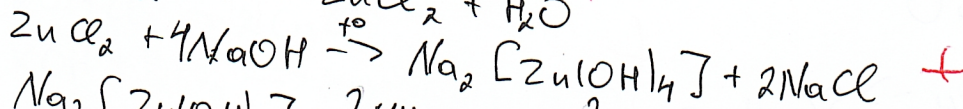
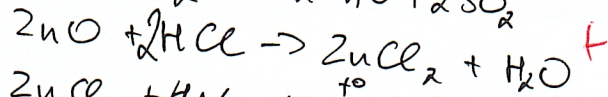
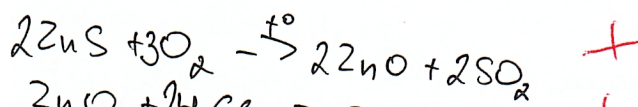
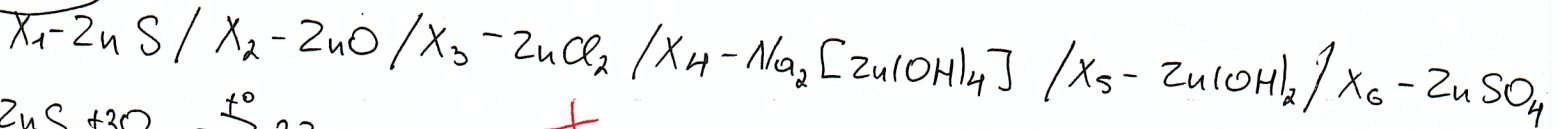
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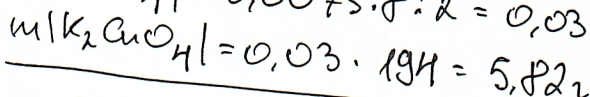
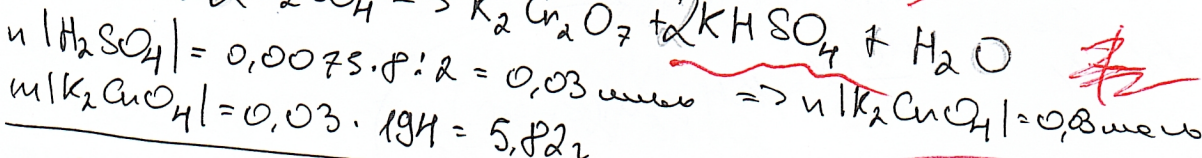
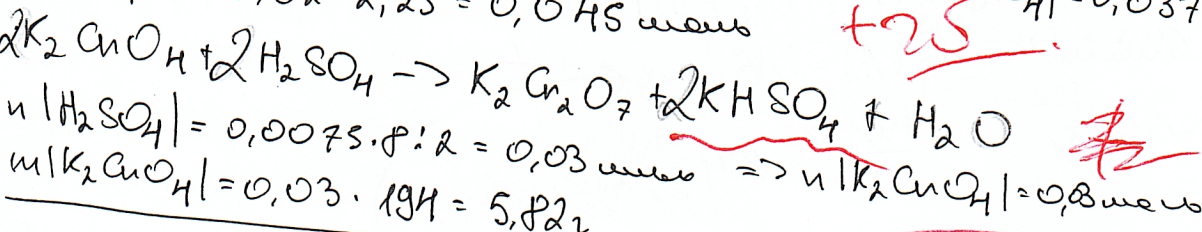
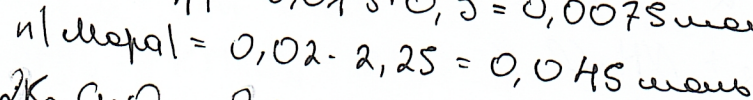
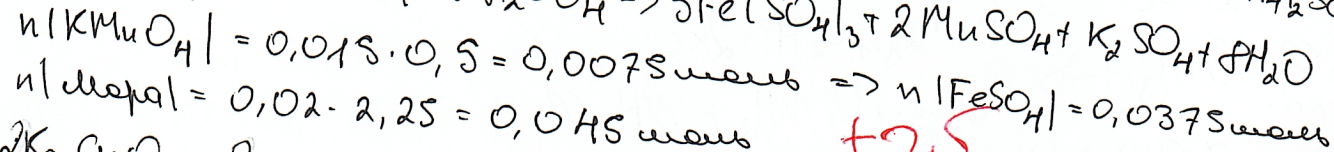
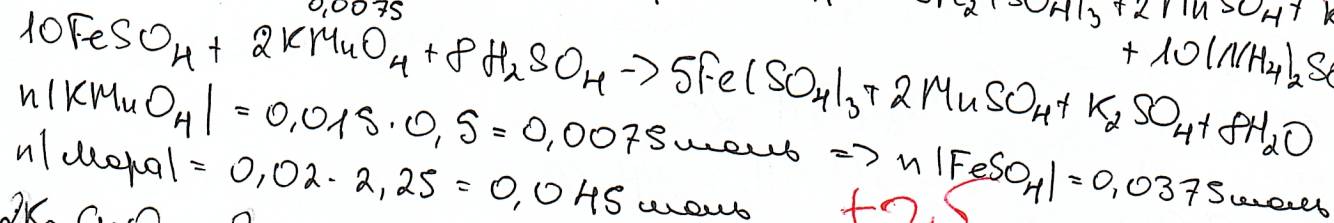
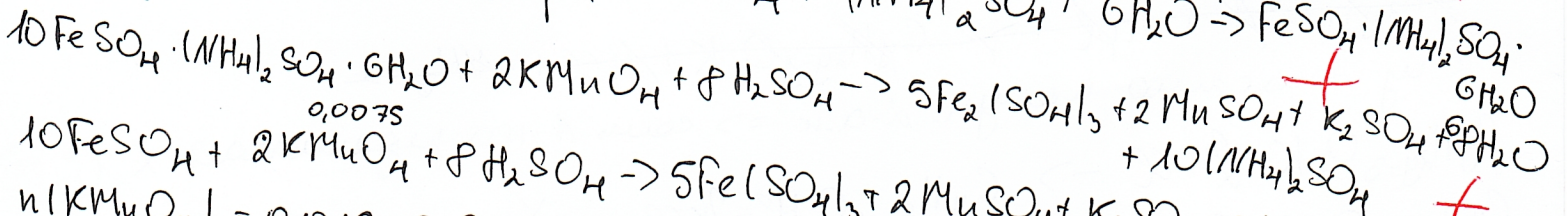
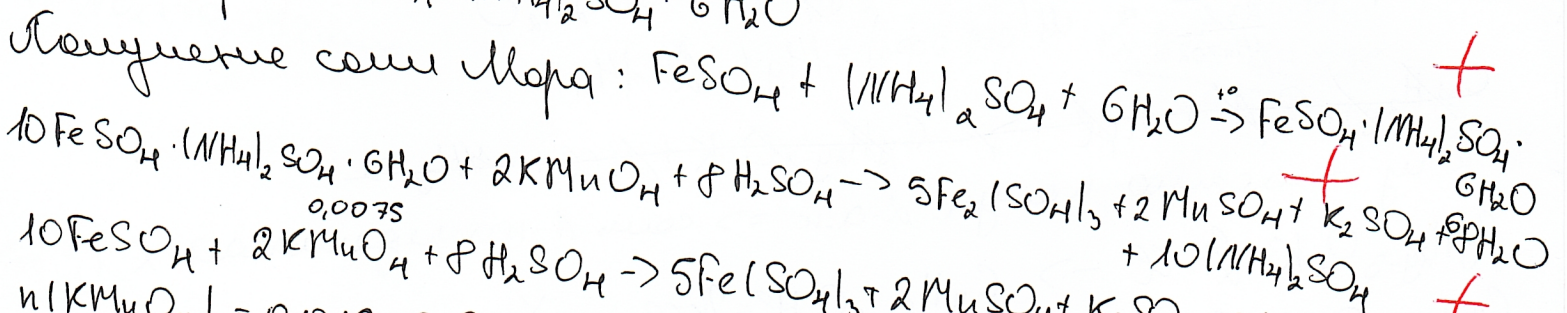
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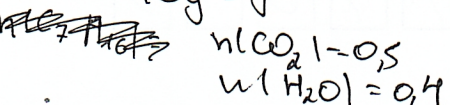
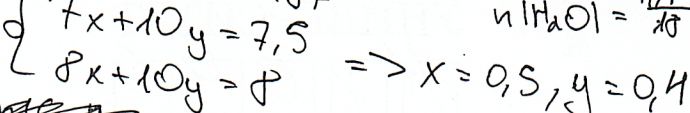
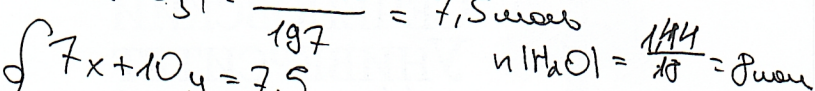
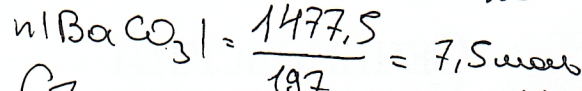
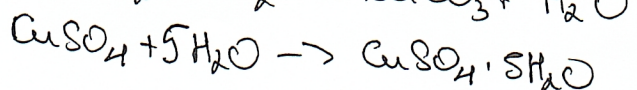
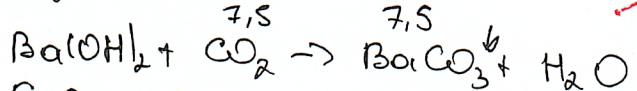
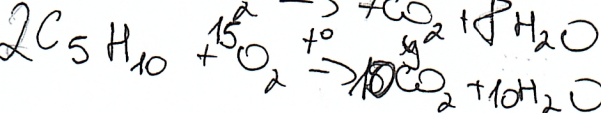
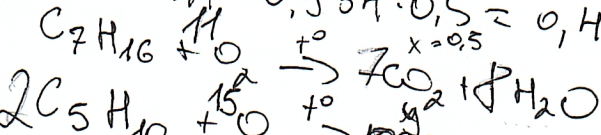
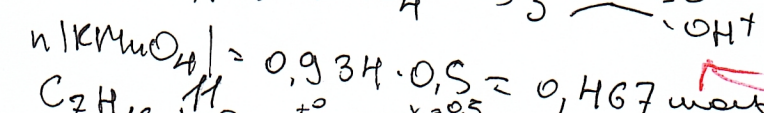
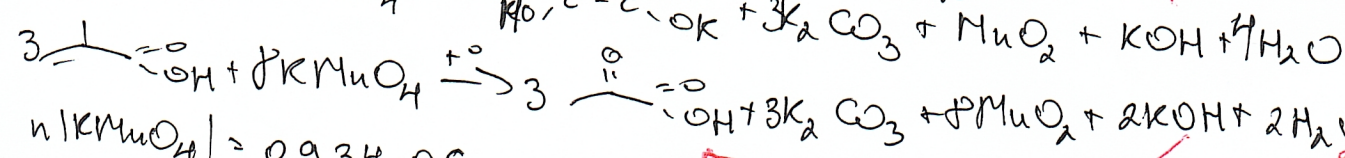
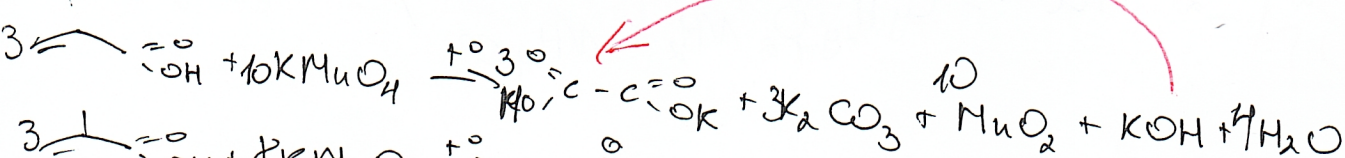


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соль Морра: $FeSO_4 \cdot (1/11H_4)_2SO_4 \cdot 6H_2O$



~3.2



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