

N1 $M(MgZn_2(PO_4)_2 \cdot 4H_2O) = 416 \text{ г/моль}$

$M(ZnO) = 81 \text{ г/моль}$

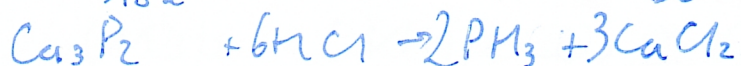
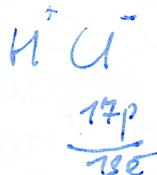
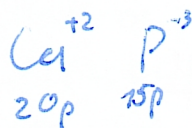
1) $\omega(Zn \text{ в } MgZn_2(PO_4)_2 \cdot 4H_2O) = \frac{130}{416} = 0,3125$

2) $\omega(Zn \text{ в } ZnO) = \frac{65}{81} = 0,8025$

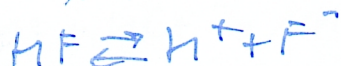
3) $\omega(Zn) = 0,3125 \cdot 0,8 + 0,8025 \cdot 0,15 = 0,3704 \Rightarrow 37,04\%$

Ответ: (37%)

N2



N4



1) $\omega = C \cdot V$

$\omega(HF) = 1,55 \cdot 0,1 = 0,155 \text{ моль}$

$\omega(HF_{\text{дис}} + H^+ + F^-) = \frac{1,014 \cdot 10^{23}}{6,02 \cdot 10^{23}} = 0,169 \text{ моль}$

x моль - протиссоциировало HF

$(0,155 - x) + 2x = 0,169$

$x = 0,0139 \approx 0,014$

$\alpha = \frac{0,014}{0,155} \cdot 100\% = 9$

Ответ: (9)

1	2	3	4	5	6	7	8	9	10
6	6	1	8	10	6	9	12	4	-



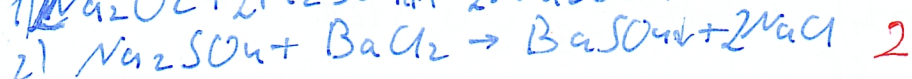
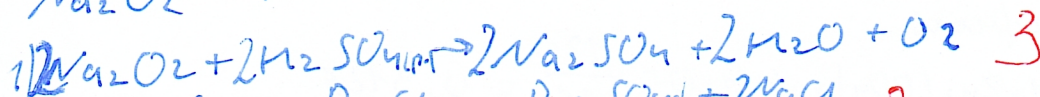
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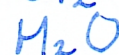
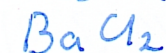
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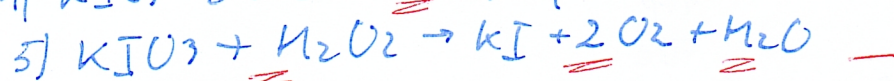
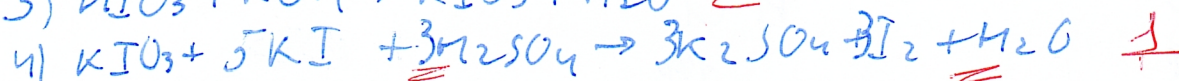
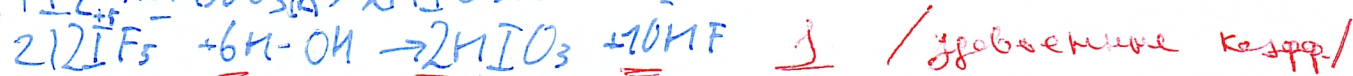
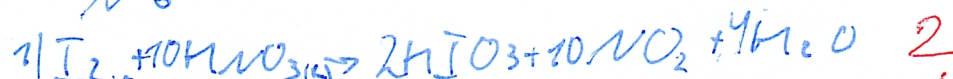
№ 5



Дано:



№ 6



№ 4

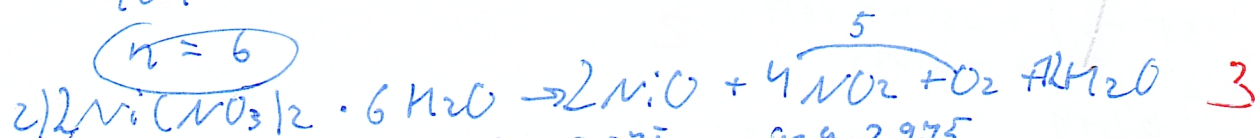


$$1) 16(6+n) = 6,857 \cdot (14 \cdot 2)$$

$$96 + 16n = 192$$

$$16n = 96$$

$$n = 6$$



$$V(\text{NO}_2 + \text{O}_2) = \frac{PV}{RT} = \frac{101,3 \cdot 9,075}{8,31 \cdot (22+273)} = \frac{919,2975}{2951,45} = 0,375 \text{ моль}$$

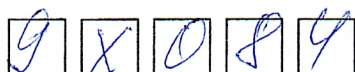
$$V(\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}) = \frac{0,375 \cdot 2 \cdot 291}{5} = 0,15 \text{ моль}$$

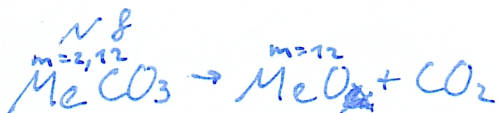
$$m = M \cdot V = 291 \cdot 0,15 = 43,65 \text{ г}$$

Ответ: 43,65 г



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+ 3

1) $m(\text{CO}_2) = 2,1 - 1 = 1,12$

2) $\nu(\text{CO}_2) = \frac{1,1}{44} = 0,025 = \nu(\text{MeO})$ +

3) $M(\text{MeO}) = \frac{m}{\nu} = \frac{1}{0,025} = 40 \text{ г/моль}$

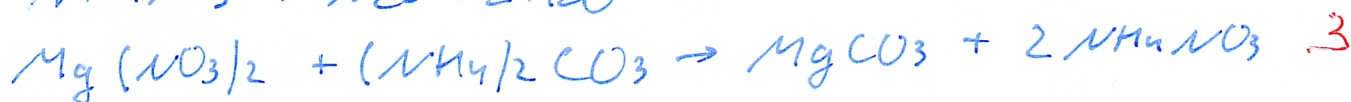
$M(\text{Me}) = 40 - M(\text{O}) = 40 - 16 = 24 \text{ г/моль}$ + 3

1) $\nu(\text{H}_2\text{O}) = \frac{1,8}{18} = 0,1 \text{ моль}$

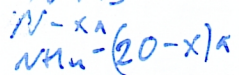
2) $\nu(\text{ZnO}) = \frac{0,1}{2} = 0,05 \text{ моль}$

3) $m(\text{ZnO}) = 4 - 1,8 = 2,2$

4) $M(\text{ZnO}) = \frac{2,2}{0,05} = 44 \text{ г/моль} \Rightarrow \text{ZnO}$ 28.2/моль 16 г/моль 3



N3



N9

