

A manganese plate was immersed in 50 g of 1,89 % solution of copper (II) chloride. The mass of manganese participated in the reaction equals

A.	945	
B.	448	
C.	265	
D.	385	
E.	224	
F.	64	

True D

A zinc plate was immersed in 15 g of 6,62 % solution of lead (II) nitrate. Mass of zinc (in mg) participated in the reaction equals

G.	993	
H.	227	
I.	414	
J.	621	
K.	65	
L.	195	

True F

An iron plate was immersed in 10 g of 3,25% of mercury (II) nitrate. Mass of mercury (in mg) precipitated on the plate equals

M.	325	
N.	308	
O.	650	
P.	402	
Q.	201	
R.	112	

True E

An iron plate was immersed in 20 g of 1,6 % solution of copper (II) sulfate. Mass of iron (in mg) participated in the reaction equals

S.	320	
T.	125	
U.	64	
V.	640	
W.	112	
X.	256	

True E

Balance the equation: $\text{NH}_3 + \text{Cl}_2 \rightarrow \text{N}_2 + \text{NH}_4\text{Cl}$ Choose coefficients for salt, ammonia and elementary substance (air component), correspondingly		
Y.	1	
Z.	3	
AA.	4	
BB.	6	
CC.	7	
DD.	8	

True A,D,F

Balance the equation: $\text{NO}_2 + \text{H}_2 \rightarrow \text{NH}_3 + \text{H}_2\text{O}$ Choose coefficients for ammonia, elementary substance, water correspondingly		
EE.	2	
FF.	3	
GG.	4	
HH.	5	
II.	6	
JJ.	7	

True F,C,A

Excess of acetic acid was treated with 89,6 l of chlorine (normal conditions). Mass of formed chloroacetic acid equals		
KK.	896 g	
LL.	94,5 g	
MM.	189 g	
NN.	378 g	
OO.	567 g	
PP.	47,25 g	

True D

HNO <sub>3</sub> can be classified as		
QQ.	oxide	
RR.	base	
SS.	acid	
TT.	salt	

True C

In the complete ionic equation for the reaction between Al(OH) <sub>3</sub> and HCl, the sum of all coefficients equals		
UU.	11	
VV.	12	
WW.	13	
XX.	14	

True D

In the molecular chemical equation for the reaction between  $\text{CaCl}_2$  and  $\text{AgNO}_3$ , the sum of all coefficients equals

YY.	3	
ZZ.	4	
AAA.	5	
BBB.	6	

True D

In the net ionic equation for the reaction between  $\text{CaCl}_2$  and  $\text{AgNO}_3$ , the sum of all coefficients equals

CCC.	3	
DDD.	4	
EEE.	5	
FFF.	6	

True A

Which atom contains twenty three electrons?

GGG.	Manganese	
HHH.	Calcium	
III.	Vanadium	
JJJ.	Cobalt	

True C

Which gas will form on dissolving of aluminium sulfide in water?

KKK.	S	
LLL.	$\text{H}_2\text{S}$	
MMM.	$\text{SO}_2$	
NNN.	$\text{H}_2\text{SO}_4$	

True B

Which of the following is a phenol?

OOO.	$\text{CH}_3\text{Cl}$	
PPP.	$\text{CH}_3\text{OH}$	
QQQ.	$\text{C}_6\text{H}_5\text{OH}$	
RRR.	$\text{C}_2\text{H}_5\text{Cl}$	

True C

Which of the following is a phenol?

SSS.	$\text{CH}_3\text{Cl}$	
TTT.	$\text{CH}_3\text{OH}$	
UUU.	$\text{C}_6\text{H}_5\text{OH}$	
VVV.	$\text{C}_2\text{H}_5\text{Cl}$	

True C

Which of the following is a precipitation reaction?	
WWW.	$\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
XXX.	$\text{Mg}(\text{OH})_2 \rightarrow \text{MgO} + \text{H}_2\text{O}$
YYY.	$\text{MgSO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + \text{MgCl}_2$
ZZZ.	$\text{MgI}_2 + \text{Br}_2 \rightarrow \text{MgBr}_2 + \text{I}_2$

True C

Which of the following is an ether?	
AAAA.	$\text{C}_2\text{H}_4$
BBBB.	$\text{CH}_4$
CCCC.	$(-\text{CH}_2-\text{CH}_2-)$
DDDD.	$\text{CH}_3-\text{O}-\text{CH}_3$

True D

Which of the following is an unsaturated hydrocarbon?	
EEEE.	$\text{C}_2\text{H}_4$
FFFF.	$\text{CH}_4$
GGGG.	$(-\text{CH}_2-\text{CH}_2-)$
HHHH.	$\text{CH}_3-\text{O}-\text{CH}_3$

True A

Which of the following will form on dissolving of sulfur (VI) oxide in water?	
IIII.	S
JJJJ.	$\text{H}_2\text{S}$
KKKK.	$\text{SO}_2$
LLLL.	$\text{H}_2\text{SO}_4$

True D

Which of the following will form on dissolving of phosphorus (V) oxide in water?	
MMMM.	$\text{H}_3\text{PO}_4$
NNNN.	P(red)
OOOO.	P(white)
PPPP.	$\text{PH}_3$

True A

Which of the following will not participate in the hydrolysis reaction?	
QQQQ.	glucose
RRRR.	benzene
SSSS.	methyl acetate
TTTT.	acetone
UUUU.	ethyl acetate
VVVV.	methyl formiate

True A,B,D

Which of the following will not react with both hydrochloric acid and sodium hydroxide?	
WWWV	methyl benzene
XXXX.	nitrophenol
YYYY.	aminoacetic acid
ZZZZ.	methylamine
AAAAA	aminopropionic acid
BBBBE	aminobutyric acid

True A

Which of the following will participate in the hydrolysis reaction?	
CCCCC	glucose
DDDDI	benzene
EEEEE	methyl acetate
FFFFF	acetone
GGGG	ethyl acetate
HHHHI	methyl formiate

True C,E,F

Which of the following will react with copper (II) hydroxide without heating?	
IIIII.	polyethylene
JJJJJ.	glycerol
KKKKK	ethylene glycol
LLLLL.	acetone
MMMM	glucose
NNNNI	benzene

True B,C,E

Which of the following will react with both hydrochloric acid and sodium hydroxide?	
OOOO	methyl benzene
PPPPF	nitrophenol
QQQQ	aminoacetic acid
RRRRF	methylamine
SSSSS	aminopropionic acid
TTTTT	aminobutyric acid

True C,E,F